

A Mountaineering Strategy to Excited States: Highly Accurate Reference Energies and Benchmarks

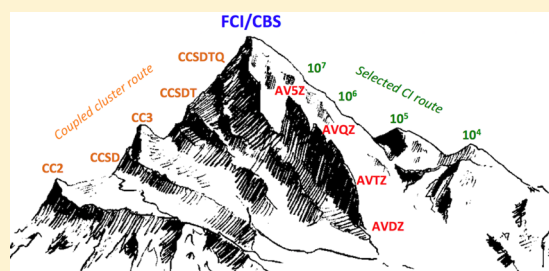
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Supporting Information

ABSTRACT: Striving to define very accurate vertical transition energies, we perform both high-level coupled cluster (CC) calculations (up to CCSDTQP) and selected configuration interaction (sCI) calculations (up to several millions of determinants) for 18 small compounds (water, hydrogen sulfide, ammonia, hydrogen chloride, dinitrogen, carbon monoxide, acetylene, ethylene, formaldehyde, methanimine, thioformaldehyde, acetaldehyde, cyclopropene, diazomethane, formamide, ketene, nitrosomethane, and the smallest streptocyanine). By systematically increasing the order of the CC expansion, the number of determinants in the CI expansion as well as the size of the one-electron basis set, we have been able to reach near full CI (FCI) quality transition energies. These calculations are carried out on CC3/*aug-cc-pVTZ* geometries, using a series of increasingly large atomic basis sets systematically including diffuse functions. In this way, we define a list of 110 transition energies for states of various characters (valence, Rydberg, $n \rightarrow \pi^*$, $\pi \rightarrow \pi^*$, singlet, triplet, etc.) to be used as references for further calculations. Benchmark transition energies are provided at the *aug-cc-pVTZ* level as well as with additional basis set corrections, in order to obtain results close to the complete basis set limit. These reference data are used to benchmark a series of 12 excited-state wave function methods accounting for double and triple contributions, namely ADC(2), ADC(3), CIS(D), CIS(D_∞), CC2, STEOM-CCSD, CCSD, CCSDR(3), CCSDT-3, CC3, CCSDT., and CCSDTQ. It turns out that CCSDTQ yields a negligible difference with the extrapolated CI values with a mean absolute error as small as 0.01 eV, whereas the coupled cluster approaches including iterative triples are also very accurate (mean absolute error of 0.03 eV). Consequently, CCSDT-3 and CC3 can be used to define reliable benchmarks. This observation does not hold for ADC(3) that delivers quite large errors for this set of small compounds, with a clear tendency to overcorrect its second-order version, ADC(2). Finally, we discuss the possibility to use basis set extrapolation approaches so as to tackle more easily larger compounds.



1. INTRODUCTION

Defining an effective method reliably providing accurate excited-state energies and properties remains a major challenge in theoretical chemistry. For practical applications, the most popular approaches are the complete active space self-consistent field (CASSCF)^{1,2} and the time-dependent density functional theory (TD-DFT)^{3,4} methods for systems dominated by static and dynamic electron correlation effects, respectively. When these schemes are not sufficiently accurate, one often uses methods including second-order perturbative corrections. For CASSCF, a natural choice is CASPT2,⁵ but this method rapidly becomes impractical for large compounds. If a single-reference method is sufficient, the most popular second-order approaches are probably the second-order algebraic diagrammatic construction, ADC(2),⁶ and the second-order coupled cluster, CC2, methods,^{7,8} that both offer an attractive $O(N^5)$ scaling (where N is the number of basis functions) allowing applications up to systems compris-

ing ca. 100 atoms. Compared to TD-DFT,⁹ these approaches have the indisputable advantage of being free of the choice of a specific exchange-correlation functional. Using ADC(2) or CC2 generally provides more systematic errors with respect to reference values than TD-DFT, although the improvements in terms of error magnitude are often rather moderate (at least for valence singlet states).^{10–12} Importantly, both ADC(n) and CC n offer a systematic pathway for improvement via an increase of the expansion order n . For example, using CCSD, CCSDT, CCSDTQ, etc., allows to check the quality of the obtained estimates. However, in practice, one can only contemplate such systematic approach and the ultimate choice of a method for excited-state calculations is often guided by previous benchmarks. These benchmark studies are either performed using experimental or theoretical reference values. While the former approach allows in principle to rely on an

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almost infinite pool of reference data, most measurements are performed in solution and provide absorption bands that can be compared to theory only with the use of extra approximations for modeling environmental and vibronic effects. In addition, the most accurate experimental data are obtained for 0–0 energies, whereas obtaining trustworthy experimental estimates of vertical transition energies is an extremely difficult task, generally requiring to back-transform spectroscopic vibronic data through a numerical process,¹³ an approach that is typically only applicable to diatomics. Consequently, it is easier to use first-principle reference values as benchmarks, as they allow to assess theoretical methods more consistently (vertical values, same geometries, no environmental effects, etc.). This is well illustrated by the recent contribution of Schwabe and Goerigk,¹⁴ who decided to compute third-order response CC (CC3)^{15,16} reference values instead of using the previously collected experimental values for the test set originally proposed by Gordon's group.¹⁷

While many benchmark sets have been proposed for excited states,^{10,11,17–29} the most praised database of theoretical excited state energies is undoubtedly the one set up by Thiel and his co-workers. In 2008, they proposed a large set of theoretical best estimates (TBE) for 28 small and medium CNOH organic compounds.³⁰ More precisely, using some literature values but mainly their own CC3/TZVP and CASPT2/TZVP results computed on MP2/6-31G(d) geometries, these authors determined 104 singlet and 63 triplet reference excitation energies. The same group soon proposed *aug-cc-pVTZ* TBE for the same set of compounds,^{31,32} though some CC3/*aug-cc-pVTZ* reference values were estimated by a basis set extrapolation technique. In their conclusion, they stated that they “expect this benchmark set to be useful for validation and development purposes, and anticipate future improvements and extensions of this set through further high-level calculations”.³⁰ The first prediction was soon realized. Indeed, both the TZVP and *aug-cc-pVTZ* TBE were applied to benchmark various computationally effective methods, including semiempirical approaches,^{33–35} TD-DFT,^{24,25,36–46} the second-order polarization propagator approximation (SOPPA),⁴⁷ ADC(2),⁴⁸ the second order *N*-electron valence perturbation theory (NEVPT2),⁴⁹ the random phase approximation (RPA),⁵⁰ as well as several CC variants.^{51–56} In contrast, even a decade after the original work appeared, the progresses aiming at improving and/or extending Thiel's set have been much less numerous. To the best of our knowledge, these extensions are limited to the more compact TZVP basis set,^{48,52,57,58} but in one case.⁵⁹ This diffuse-less basis set offers clear computational advantages and avoids some state mixing. However, it has a clear tendency to overestimate transition energies, especially for Rydberg states, and it makes comparisons between methods more difficult as basis set dependencies are significantly different in wave function-based and density-based methods.⁶⁰

Let us now briefly review these efforts. In 2013, Watson et al. obtained with the TZVP basis set and CCSDT-3—a method employing an iterative approximation of the triples — transition energies very similar to the CC3 values.⁵⁷ Nevertheless, as noted the same year by Nooijen and co-workers who also reported CCSDT-3/TZVP values,⁵² “the relative accuracy of EOM-CCSDT-3 versus CC3 compared to full CI (or EOM-CCSDT) is not well established”. In 2014, Dreuw and co-workers performed ADC(3) calculations on Thiel's set and concluded that “based on the quality of the existing benchmark set

it is practically not possible to judge whether ADC(3) or CC3 is more accurate”. The same year, Kannar and Szalay, revisited Thiel's set and proposed CCSDT/TZVP reference energies for 17 singlet states of six molecules.⁵⁸ Recently the same group reported CCSDT/*aug-cc-pVTZ* transition energies for valence and Rydberg states of five compact molecules,⁵⁹ and used these values to benchmark several simpler CC approaches. To the best of our knowledge, these stand as the highest-level values reported to date. However, it remains difficult to know if these CCSDT transition energies are significantly more accurate than their CC3, CCSDT-3 or ADC(3) counterparts. Indeed, for the $\pi \rightarrow \pi^*$ valence singlet excited state of ethylene, the CC3/TZVP, CCSDT/TZVP and CCSDTQ/TZVP estimates of 8.37, 8.38, and 8.36 eV (respectively) are nearly identical.⁵⁸

Herein, we propose to continue the quest for ultra-accurate excited-state reference energies. First, although this prevents direct comparisons with previously published data, we decided to use more accurate CC3/*aug-cc-pVTZ* geometries for all the compounds considered here. Second, we employ only diffuse-containing Dunning basis sets to be reasonably close from the complete basis set limit. Third, we climb the mountain via two faces following: (i) the CC route (up to the highest computationally possible order) and (ii) the configuration interaction (CI) route with the help of selected CI (sCI) methods. By comparing the results of these two approaches, it is possible to get some reliable information about how far our results are from the full CI (FCI) ones. Fourth, in order not to limit our investigation to vertical absorption, we also report, in a few cases, fluorescence energies. Of course, such extreme choices impose drastic restrictions on the size of the molecules one can treat. However, we claim here that they allow to accurately estimate the FCI result for most excited states.

2. COMPUTATIONAL DETAILS

2.1. Geometries. All geometries are obtained at the CC3/*aug-cc-pVTZ* level without applying the frozen core approximation. These geometries are available in the [Supporting Information](#). While several structures are extracted from ref 61 (acetylene, diazomethane, ethylene, formaldehyde, ketene, nitrosomethane, thioformaldehyde and streptocyanine-C1), additional optimizations are performed here following the same protocol as in that earlier work. First, we optimize the structures and compute the vibrational spectra at the CCSD/def2-TZVPP level⁶² with Gaussian16.⁶³ These calculations confirm the minima nature of the obtained geometries.⁶⁴ We then reoptimize the structures at the CC3/*aug-cc-pVTZ* level^{15,16} using Dalton⁶⁵ and/or CFOUR,⁶⁶ depending on the size and symmetry of the molecule. CFOUR advantageously provides analytical CC3 gradients for ground-state structures. For the CCSD calculations, the energy and geometry convergence thresholds are systematically tightened to 10^{-10} – 10^{-11} au for the SCF energy, 10^{-8} – 10^{-9} au for the CCSD energy, and 10^{-7} – 10^{-8} au for the EOM-CCSD energy in the case of excited-state geometry optimizations. To check that the structures correspond to genuine minima, the (EOM-)CCSD gradients are differentiated numerically to obtain the vibrational frequencies. The CC3 optimizations are performed with the default convergence thresholds of Dalton or CFOUR without applying the frozen core approximation.

2.2. Coupled Cluster Calculations. Unless otherwise stated, the CC transition energies⁶⁷ are computed in the frozen-core approximation (large cores for CI and S). We use several codes to achieve our objectives, namely CFOUR,⁶⁶

Table 1. Vertical Transition Energies for the Three Lowest Singlet and Three Lowest Triplet Excited States of Water (Top), the Four Lowest Singlet and the Lowest Triplet States of Ammonia (Center), and the Lowest Singlet State of Hydrogen Chloride (Bottom)[†]

state	water												lit.		
	<i>aug-cc-pVDZ</i>					<i>aug-cc-pVTZ</i>				<i>aug-cc-pVQZ</i>			exp. ^a	th. ^b	th. ^c
	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI			
¹ B ₁ (n → 3s)	7.51	7.50	7.53	7.53	7.53	7.60	7.59	7.62	7.62	7.65	7.64	7.68	7.41	7.81	7.57
¹ A ₂ (n → 3p)	9.29	9.28	9.31	9.32	9.32	9.38	9.37	9.40	9.41	9.43	9.41	9.46	9.20	9.30	9.33
¹ A ₁ (n → 3s)	9.92	9.90	9.94	9.94	9.94	9.97	9.95	9.98	9.99	10.00	9.98	10.02	9.67	9.91	9.91
³ B ₁ (n → 3s)	7.13	7.11	7.14	7.14	7.14	7.23	7.22	7.24	7.25	7.28	7.26	7.30	7.20	7.42	7.21
³ A ₂ (n → 3p)	9.12	9.11	9.14	9.14	9.14	9.22	9.20	9.23	9.24	9.26	9.25	9.28	8.90	9.42	9.19
³ A ₁ (n → 3s)	9.47	9.45	9.48	9.49	9.49	9.52	9.50	9.53	9.54	9.56	9.54	9.58	9.46	9.78	9.50
state	hydrogen sulfide												lit.		
	<i>aug-cc-pVDZ</i>					<i>aug-cc-pVTZ</i>				<i>aug-cc-pVQZ</i>			exp. ^d	exp. ^e	th. ^f
	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI			
¹ A ₂ (n → 4p)	6.29	6.29	6.29	6.29	6.29	6.19	6.18	6.18	6.18	6.16	6.15	6.15			6.12
¹ B ₁ (n → 4s)	6.10	6.10	6.10	6.10	6.10	6.24	6.24	6.24	6.24	6.29	6.29	6.29	6.33		6.27
³ A ₂ (n → 4p)	5.91	5.90	5.90	5.90	5.90	5.82	5.81	5.81	5.81	5.80	5.79	5.79		5.8	5.78
³ B ₁ (n → 4s)	5.75	5.75	5.75	5.75	5.75	5.88	5.88	5.88	5.89	5.93	5.93	5.93		5.4	5.92
state	ammonia												lit.		
	<i>aug-cc-pVDZ</i>					<i>aug-cc-pVTZ</i>				<i>aug-cc-pVQZ</i>			exp. ^g	exp. ^h	th. ⁱ
	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI			
¹ A ₂ (n → 3s)	6.46	6.46	6.48	6.48	6.48	6.57	6.57	6.59	6.59	6.61	6.61	6.64	6.38	6.39	6.48
¹ E(n → 3p)	8.06	8.06	8.08	8.08	8.08	8.15	8.14	8.16	8.16	8.18	8.17	8.22	7.90	7.93	8.02
¹ A ₁ (n → 3p)	9.66	9.66	9.68	9.68	9.68	9.32	9.31		9.33	9.11	9.10	9.14	8.14	8.26	8.50
¹ A ₂ (n → 4s)	10.40	10.39	10.41	10.41	10.41	9.95	9.94		9.96	9.77	9.77				9.03
³ A ₂ (n → 3s)	6.18	6.18	6.19	6.19	6.19	6.29	6.29	6.30	6.31	6.33	6.33	6.35	6.02 ^j		
state	hydrogen chloride												lit.		
	<i>aug-cc-pVDZ</i>					<i>aug-cc-pVTZ</i>				<i>aug-cc-pVQZ</i>			th. ^k		
	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI			
¹ Π (CT)	7.82	7.81	7.82	7.82	7.82	7.84	7.83	7.84	7.84	7.89	7.88 ^l	7.88		8.23	

^aEnergy loss experiment from ref 98. ^bMRCI+Q/*aug-cc-pVTZ* calculations from ref 99. ^cMRCC/*aug-cc-pVTZ* calculations from ref 100. ^dVUV experiment from ref 101. ^eElectron impact experiment from ref 102. ^fCASPT2/*d-aug-cc-pVQZ* results from ref 103. ^gElectron impact experiment from ref 104. ^hElectron impact experiment from ref 105. ⁱEOM-CCSD(\bar{T})/*aug-cc-pVTZ* with extra *diffuse* calculations from ref 106. ^jDeduced from the 6.38 eV value of the ¹A₂(n → 3s) state and the −0.36 eV shift reported for the 0–0 energies compared to the corresponding singlet state in ref 107, a splitting consistent with an earlier estimate of −0.39 eV given in ref 108. ^kCC2/*cc-pVTZ* from ref 22.; ^lThe CCSDTQ/*aug-cc-pVQZ* value is 7.88 eV as well. ^mAll states of water and ammonia have a Rydberg character, whereas the lowest state of hydrogen chloride is a charge-transfer state. All values are in eV.

Dalton,⁶⁵ Gaussian16,⁶³ Orca,⁶⁸ MRCC,^{69,70} and Q-Chem.⁷¹ Globally, we use CFOUR for both CCSDT-3^{72,73} and CCSDT⁷⁴ calculations, Dalton to perform the CIS(D),^{75,76} CC2,^{7,8} CCSD,⁶² CCSDR(3),⁷⁷ and CC3^{15,16} calculations, Gaussian for the CIS(D)^{75,76} and CCSD,⁶² Orca for the similarity-transformed EOM-CCSD (STEOM-CCSD)^{56,78} calculations, Q-Chem for ADC(2) and ADC(3) calculations, and MRCC for the CIS(D_∞),⁷⁹ CCSDT,⁷⁴ and CCSDTQ⁸⁰ (and higher) calculations. As we mainly report transition energies, it is worth noting that the linear-response (LR) and equation-of-motion (EOM) formalisms provide identical results. Nevertheless, the oscillator strengths characterizing the excited states are obtained at the (LR) CC3 level with Dalton. Default program setting are generally applied, and when modified they are tightened. For the STEOM-CCSD calculations which relies on natural transition orbitals, it was checked that each state is characterized by an active character percentage of 98% or larger (states not matching this criterion are not reported). Nevertheless, the obtained results do slightly depend on the number of states included in the calculations, and we found typical variations of ±0.01–0.05 eV. For all

calculations, we use the well-known Dunning's *aug-cc-pVXZ* (X = D, T, Q and 5) atomic basis sets, as well as some doubly- and triply augmented basis sets of the same series (*d-aug-cc-pVXZ* and *t-aug-cc-pVXZ*).

2.3. Selected Configuration Interaction Methods.

Alternatively to CC, we also compute transition energies using a selected CI (sCI) approach, an idea that goes back to 1969 in the pioneering works of Bender and Davidson,⁸¹ and Whitten and Hackmeyer.⁸² Recently, sCI methods have demonstrated their ability to reach near FCI quality energies for small organic and transition metal-containing molecules.^{83–92} To avoid the exponential increase of the size of the CI expansion, we employ the sCI algorithm CIPSI^{83,93,94} (Configuration Interaction using a Perturbative Selection made Iteratively) to retain only the energetically relevant determinants. To do so, the CIPSI algorithm uses a second-order energetic criterion to select perturbatively determinants in the FCI space.^{83,85,87,92} In the numerical examples presented below, our CI expansions contain typically about a few millions of determinants. We refer the interested readers to refs 92 and

95 for more details about the general philosophy of sCI methods.

In order to treat the electronic states of a given spin manifold on equal footing, a common set of determinants is used for all states. Moreover, to speed up convergence to the FCI limit, a common set of natural orbitals issued from a preliminary (smaller) sCI calculation is employed. All sCI calculations have been performed in the frozen-core approximation. For a given basis set, we estimate the FCI limit using the approach introduced recently by Holmes et al.⁹⁰ in the context of the (selected) heat-bath CI method, and used with success, even for challenging chemical situations.^{89,91,92} More precisely, we linearly extrapolate the sCI energy E_{sCI} as a function of E_{PT2} , which is an estimate of the truncation error in the sCI algorithm, i.e., $E_{\text{PT2}} \approx E_{\text{FCI}} - E_{\text{sCI}}$. When $E_{\text{PT2}} = 0$, the FCI limit has effectively been reached. Here, E_{PT2} is efficiently evaluated with a recently proposed hybrid stochastic-deterministic algorithm.⁹⁶ Note that we do not report error bars because the statistical errors originating from this algorithm are orders of magnitude smaller than the extrapolation errors. In practice, the extrapolation is based on the two largest sCI wave functions; i.e., we perform a two-point extrapolation, which is justified here because of the quasi-linear behavior of the sCI energy as a function of E_{PT2} . Estimating the extrapolation error is a complicated task with no well-defined method to do so. In practice, we have observed that this extrapolation procedure is robust and provides FCI estimates within ± 0.02 eV. When the convergence to the FCI limit is too slow to provide reliable estimates, the number of significant digits reported has been reduced accordingly. From herein, the extrapolated FCI results are simply labeled exFCI. Several illustrative examples are reported in [Supporting Information](#) where we compare different types of extrapolations for several molecules (see [Figure S1](#) and [Table S11](#)). In particular, diazomethane and streptocyanine-C1 can be considered as “difficult” cases (*vide infra*), and the results reported in [Supporting Information](#) show that, even in these challenging situations, the two-point linear extrapolation is fairly robust. Moreover, additional points do not significantly alter the exFCI estimates (typically 0.01 eV or less).

All the sCI calculations are performed with the electronic structure software QUANTUM PACKAGE, developed in Toulouse and freely available.⁹⁷ Additional information about the sCI wave functions, excitations energies as well as their extrapolated values can be found at the end of the [Supporting Information](#).

3. RESULTS AND DISCUSSION

In the discussion below, we first discuss specific molecules of increasing size and compare the results obtained with exFCI and CC approaches, starting with the CC3 method for the latter. This first part is performed applying systematically the frozen-core approximation. We next define two series of TBE, one at the frozen-core *aug-cc-pVTZ* level, and one close to complete basis set limit by applying corrections for frozen-core and basis set effects. In a following stage, we assess the performances of several popular wave function methods using the former benchmark as reference. Finally, we discuss the performances of basis set extrapolation approaches starting from a compact basis. Unless otherwise stated, we considered the exFCI values as benchmarks.

3.1. Water, Hydrogen Sulfide, Ammonia, and Hydrogen Chloride. Because of its small size and ubiquitous role in

life, water is often used as a test case for Rydberg excitations. Indeed, it is part of Head–Gordon’s,²¹ Gordon’s¹⁷ and Truhlar–Gagliardi’s²⁹ data sets of compounds, and it has been investigated at many levels of theory.^{99,100,103,109} Our results are collected in [Table 1](#). With the *aug-cc-pVDZ* basis, there is an nearly perfect agreement between the exFCI values and the transition energies obtained with the two largest CC expansions, namely CCSDTQ and CCSDTQP. Indeed, the largest discrepancy is as small as 0.01 eV, and it is therefore reasonable to state that the FCI limit has been reached with that specific basis set. Compared to the exFCI results, the CCSDT values are systematically too low, with an average error of -0.03 eV. The same trend of underestimation is found with CC3, though with smaller absolute deviations for all states. Unsurprisingly, for Rydberg states, increasing the basis set size has a significant impact, and it tends to increase the computed transition energies in water. However, this effect is very similar for all methods listed in [Table 1](#). This means that, on the one hand, the tendency of CCSDT to provide slightly too small transition energies pertains with both *aug-cc-pVTZ* and *aug-cc-pVQZ*, and, on the other hand, that estimating the basis set effect with a “cheap” method is possible. Indeed, adding to the exFCI/*aug-cc-pVDZ* energies, the difference between CC3/*aug-cc-pVQZ* and CC3/*aug-cc-pVDZ* results would deliver estimates systematically within 0.01 eV of the actual exFCI/*aug-cc-pVQZ* values. Such basis set extrapolation approach was already advocated for lower-order CC expansions,^{31,110} and it is therefore not surprising that it can be applied with refined models. As it can be seen in [Table S1](#) in the [Supporting Information](#), further extension of the basis set or correlation of the 1s electron have small impacts, except for the Rydberg 1A_1 state. Eventually, as evidenced by the data from the rightmost columns of [Table 1](#), the present estimates are in good agreement with previous MRCC values determined on the experimental geometry,¹⁰⁰ whereas the experimental values offer qualitative comparisons only, for reasons discussed elsewhere.⁹⁸ We underline that some of the 2013 measurements reported in [Table 1](#) significantly differ from previous electron impact data,¹¹¹ that were used previously as references,¹⁷ with, e.g., a 0.2 eV discrepancy between the two experiments for the lowest triplet state.

As water, hydrogen sulfide was also the subject of several high-level theoretical investigations,^{103,112–114} which are necessary because there are either no (lowest 1A_2 state) or only a few experimental data available for the Rydberg states of H_2S ,^{101,102,115,116} especially as no accurate value could be measured for the first 1A_2 state. As can be seen in [Table 1](#), for a given basis set all tested CC methods provide very similar results, systematically within 0.01 eV of the exFCI results. In contrast, the basis set has a significant impact, e.g., the two lowest singlet states switch order when going from *aug-cc-pVDZ* to *aug-cc-pVTZ* and the same is true for the two lowest triplet states. Our results are also very consistent with the CASPT2/*d-aug-cc-pVQZ* values given in ref 103, confirming that a near FCI limit has been reached.

Ammonia is also another popular molecule for evaluating Rydberg excitations, and it was previously investigated at several levels of theory.^{14,21,106,117} As in the case of water, we note a nearly perfect match between the CCSDTQ and exFCI estimates with both the *aug-cc-pVDZ* and *aug-cc-pVTZ* atomic basis sets, indicating that the FCI limit is reached. Both CC3 and CCSDT are close to this limit, and the former model slightly outperforms the latter. For ammonia, the basis set

Table 2. Vertical Transition Energies for Various Excited States of Dinitrogen (Top) and Carbon Monoxide (Bottom)^a

state	dinitrogen										litt.				
	aug-cc-pVDZ					aug-cc-pVTZ					aug-cc-pVQZ		exp. ^b	th. ^c	
	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	exFCI	exp. ^b	th. ^c
¹ Π _g (n → π*)	9.44	9.41	9.41	9.41	9.41	9.34	9.33	9.32	9.34	9.34	9.33	9.31	9.34	9.31	9.27
¹ Σ _g ⁺ (π → π*)	10.06	10.06	10.06	10.05	10.05	9.88	9.89	9.88	9.88	9.88	9.87	9.88	9.92	9.92	10.09
¹ Δ _g (π → π*)	10.43	10.44	10.43	10.43	10.43	10.29	10.30	10.29	10.29	10.29	10.27	10.28	10.31	10.27	10.54
¹ Σ _g ⁺ (R)	13.23	13.20	13.18	13.18	13.18	13.01	13.00	12.97	12.98	12.98	12.90	12.89	12.89	12.2	12.20
¹ Π _u (R)	13.28	13.17	13.13	13.13	13.12	13.22	13.14	13.09	13.03	13.03	13.17	13.1 ^d	13.1 ^d	12.78	12.90
¹ Σ _g ⁺ (R)	13.14	13.13	13.11	13.11	13.11	13.12	13.12	13.09	13.09	13.09	13.09	13.09	13.2 ^d	12.96	12.82
¹ Π _u (R)	13.64	13.59	13.56	13.56	13.56	13.49	13.45	13.42	13.46	13.46	13.42	13.37	13.7 ^d	13.10	13.61
³ Σ _g ⁺ (π → π*)	7.67	7.68	7.69	7.70	7.70	7.68	7.69	7.70	7.70	7.70	7.71	7.71	7.74	7.75	7.56
³ Π _g (n → π*)	8.07	8.06	8.05	8.05	8.05	8.04	8.03	8.02	8.01	8.04	8.04	8.04	8.03	8.04	8.05
³ Δ _u (π → π*)	8.97	8.96	8.96	8.96	8.96	8.87	8.87	8.87	8.87	8.87	8.87	8.87	8.88	8.88	8.93
³ Σ _g ⁺ (π → π*)	9.78	9.76	9.75	9.75	9.75	9.68	9.68	9.66	9.66	9.66	9.68	9.66	9.66	9.67	9.86

state	carbon monoxide										litt.				
	aug-cc-pVDZ					aug-cc-pVTZ					aug-cc-pVQZ		exp. ^c	th. ^e	
	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	CCSDTQ	CCSDTQP	exFCI	CC3	CCSDT	exFCI	exp. ^c	th. ^e
¹ Π(n → π*)	8.57	8.57	8.56	8.56	8.57	8.49	8.49	8.48	8.49	8.49	8.47	8.48	8.50	8.51	8.83
¹ Σ ⁻ (π → π*)	10.12	10.06	10.06	10.06	10.05	9.99	9.94	9.93	9.92	9.92	9.99	9.94	9.99	9.88	9.97
¹ Δ(π → π*)	10.23	10.18	10.17	10.17	10.16	10.12	10.08	10.07	10.06	10.06	10.12	10.07	10.11	10.23	10.00
¹ Σ ⁺ (R)	10.92	10.94	10.93	10.92	10.94	10.94	10.99	10.96	10.95	10.95	10.90	10.95	10.96	10.78	10.98
¹ Σ ⁺ (R)	11.48	11.52	11.51	11.51	11.52	11.49	11.54	11.52	11.52	11.52	11.46	11.51	11.53	11.40	
¹ Π(R)	11.74	11.77	11.76	11.75	11.76	11.69	11.74	11.72	11.72	11.72	11.63	11.69	11.70	11.53	
³ Π(n → π*)	6.31	6.30	6.29	6.28	6.29	6.30	6.30	6.28	6.28	6.28	6.30	6.30	6.29	6.32	6.41
³ Σ ⁺ (π → π*)	8.45	8.43	8.44	8.44	8.46	8.45	8.42	8.44	8.45	8.45	8.48	8.45	8.49	8.51	8.39
³ Δ(π → π*)	9.37	9.33	9.34	9.34	9.33	9.30	9.26	9.26	9.27	9.27	9.31	9.26	9.29	9.36	9.23
³ Σ ⁻ (π → π*)	9.89				9.83	9.82			9.80	9.80	9.82	9.78	9.78	9.88	9.60
³ Σ ⁺ (R)	10.39	10.42	10.42	10.41	10.41	10.45	10.50	10.48	10.47	10.47	10.44	10.49		10.4 ^h	

^aExperimental vertical values given in ref 13 and computed from the spectroscopic constants of ref 118. ^bExperimental vertical values given in ref 119 and computed from the spectroscopic constants of ref 118. ^cMRCSSD/6-311G with one additional *d* calculations from ref 119. ^dCI convergence too slow to provide estimates to 0.01 eV. ^eExperimental vertical values given in ref 120 and computed from the spectroscopic constants of ref 118. ^fCCSDT/PVTZ+ results from ref 121. ^gCASSCF(10,10)/cc-pVTZ results from ref 122. ^hOnly one digit reported for that state, see ref 120. ⁱR stands for Rydberg states. All values are in eV.

Table 3. Vertical (Absorption) Transition Energies for the Five Lowest Low-Lying Valence Excited States of Acetylene (Top) and the Three Lowest Singlet and Triplet Excited States of Ethylene (Bottom)^g

state	acetylene								lit.		
	aug-cc-pVDZ				aug-cc-pVTZ				exp. ^a	th. ^b	th. ^c
	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI				
¹ Σ _u ⁻ (π → π*)	7.21	7.21	7.21	7.20	7.09	7.09	7.10	7.1	6.96	7.10	
¹ Δ _u (π → π*)	7.51	7.52	7.52	7.51	7.42	7.43	7.44	7.2	7.30	7.43	
³ Σ _u ⁺ (π → π*)	5.48	5.49	5.50	5.50	5.50	5.51	5.53	5.2	5.26	5.58	
³ Δ _u (π → π*)	6.46	6.46	6.46	6.46	6.40	6.39	6.40	6.0	6.20	6.41	
³ Σ _u ⁻ (π → π*)	7.13	7.14	7.14	7.14	7.07		7.08	7.1	6.90	7.05	
¹ A _u [F](π → π*)	3.70	3.72	3.70	3.71	3.64	3.66	3.64				
¹ A ₂ [F](π → π*)	3.92	3.94	3.93	3.93	3.84	3.86	3.85				

state	ethylene								lit.	
	aug-cc-pVDZ				aug-cc-pVTZ				exp. ^d	th. ^e
	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI			
¹ B _{3u} (π → 3s)	7.29	7.29	7.30	7.31	7.35	7.37	7.39	7.11	7.45	
¹ B _{1u} (π → π*)	7.94	7.94	7.93	7.93	7.91	7.92	7.93	7.60	8.00	
¹ B _{1g} (π → 3p)	7.97	7.98	7.99	8.00	8.03	8.04	8.08	7.80	8.06	
³ B _{1u} (π → π*)	4.53	4.54	4.54	4.55	4.53	4.53	4.54	4.36	4.55	
³ B _{3u} (π → 3s)	7.17	7.18	7.18	7.16	7.24	7.25	<i>f</i>	6.98	7.29	
³ B _{1g} (π → 3p)	7.93	7.94	7.94	7.93	7.98	7.99	<i>f</i>	7.79	8.02	

^aElectron impact experiment from ref 129. Note that the 7.1 eV value for the Σ_u⁻ singlet and triplet states should be viewed as a tentative assignment. ^bLS-CASPT2/*aug*-ANO calculations from ref 124. ^cMR-AQCC/extrap. calculations from ref 126. ^dExperimental values collected from various sources from ref 116. (see discussions in refs 30, 130, and 131). ^eBest composite theory from ref 131, close to FCI. ^fCI convergence too slow to provide estimates reliable to 0.01 eV. ^gFor acetylene, we also report the vertical emission (denoted [F]) obtained from the lowest *trans* and *cis* isomers. All values are in eV.

effects are particularly strong for the third and fourth singlet excited states but these basis set effects are nearly transferrable from one method to another. In fact, as hinted by the large differences between the *aug*-cc-pVTZ and *aug*-cc-pVQZ results in Table 1, these two high-lying states require the use of additional diffuse orbitals to attain convergence. The CC3/*t*-*aug*-cc-pVQZ values of 8.60 and 9.15 eV (see Table S1 in the Supporting Information), are close from the previous results of Bartlett and co-workers,¹⁰⁶ who also applied extra diffuse orbitals in their calculations relying on approximate triples (see the footnotes in Table 1). As in water, the experimental values do not provide sufficiently clear-cut results to ultimately decide which method is the most accurate. Indeed, the vertical experimental estimates reported in Table 1 differ significantly from the more trustworthy adiabatic values with variations of ca. 0.5 eV.¹⁰⁶ Consequently, a good match between an experimental measurement and a theoretical calculation determined with a compact basis set is, in the present case, a sign of lucky cancellation of errors.

Hydrogen chloride was less frequently used in previous benchmarks, but is included in Tozer's set as an example of charge-transfer (CT) state.²² Again, the results listed at the bottom of Table 1 demonstrate a remarkable consistency between the various theories. Though large frozen cores are used during the calculations, this does not strongly impact the results, as can be deduced from the data of Table S1. As expected, the absorption band corresponding to this CT state is very broad experimentally (starting at 5.5 eV and peaking at 8.1 eV),¹¹⁸ making direct comparisons tricky.

3.2. Dinitrogen and Carbon Monoxide. Dinitrogen is a simple diatomic compound for which the low-lying valence and Rydberg states have been investigated at several levels of theory.^{13,22,119,121} With a numerical solution of the nuclear

Schrödinger equation, it is possible to treat the experimental spectroscopic constants,¹¹⁸ so as to obtain reliable vertical estimates, and this procedure was applied previously.^{13,119,123} While such approach is supposedly providing experimental vertical excited-state energies with a ca. 0.01 eV error only, it remains that significant excitation energy differences have been reported for the two lowest ¹Π_u states (see Table 2). As in the previous cases, we find a remarkable agreement between the CCSDTQ and exFCI estimates for most cases in which both could be determined. The only exceptions are the two ¹Π_u states with the *aug*-cc-pVTZ basis, but in these two cases, the CC expansion is also converging more slowly than usual, which is consistent with the relatively small degree of single excitation character in these two states (82.9 and 87.4% according to CC3). In contrast to water and ammonia, CCSDT outperforms CC3 with respective mean absolute deviation (MAD) compared to exFCI of 0.02 and 0.04 eV, when using the *aug*-cc-pVDZ basis set. As it can be deduced from Table S2 in the Supporting Information, the basis set corrections are negligible for all valence states, but significant for some of the Rydberg states, especially ¹Σ_g⁺, that requires two sets of diffuse orbitals to be reasonably close from the basis set limit. Applying CC3/*d*-*aug*-cc-pVSZ corrections to the most accurate exFCI data, once can determine TBE values (*vide infra*) that deviate only by 0.02 eV on (absolute) average compared to the experimental estimates for the seven valence states of dinitrogen. Considering the expected inaccuracy of 0.01 eV of the reference values, chemical accuracy is obviously reached without any experimental input. The deviations are about twice larger for the Rydberg states. Nevertheless, for the two ¹Π_u states, our TBE values, determined on the basis of exFCI/*aug*-cc-pVTZ results are 12.73 and 13.27 eV (*vide infra*). This indicates that for the lowest ¹Π_u state the estimate of ref 13

Table 4. Vertical (Absorption) Transition Energies for Various Excited States of Formaldehyde (Top), Methanimine (Center), and Thioformaldehyde (Bottom)^h

state	formaldehyde								lit.		
	<i>aug-cc-pVDZ</i>				<i>aug-cc-pVTZ</i>				exp. ^a	th. ^b	th. ^c
	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI				
¹ A ₂ (n → π*)	4.00	3.99	4.00	3.99	3.97	3.95	3.98	4.07	3.98	3.88	
¹ B ₂ (n → 3s)	7.05	7.04	7.09	7.11	7.18	7.16	7.23	7.11	7.12		
¹ B ₂ (n → 3p)	8.02	8.00	8.04	8.04	8.07	8.07	8.13	7.97	7.94	8.11	
¹ A ₁ (n → 3p)	8.08	8.07	8.12	8.12	8.18	8.16	8.23	8.14	8.16		
¹ A ₂ (n → 3p)	8.65	8.63	8.68	8.65	8.64	8.61	8.67	8.37	8.38		
¹ B ₁ (σ → π*)	9.31	9.29	9.30	9.29	9.19	9.17	9.22		9.32	9.04	
¹ A ₁ (π → π*)	9.59	9.59	9.54	9.53	9.48	9.49	9.43		9.83	9.29	
³ A ₂ (n → π*)	3.58	3.57	3.58	3.58	3.57	3.56	3.58	3.50		3.50	
³ A ₁ (π → π*)	6.09	6.08	6.09	6.10	6.05	6.05	6.06	5.86		5.87	
³ B ₂ (n → 3s)	6.91	6.90	6.95	6.95	7.03	7.02	7.06	6.83			
³ B ₂ (n → 3p)	7.84	7.82	7.86	7.87	7.92	7.90	7.94	7.79			
³ A ₁ (n → 3p)	7.97	7.95	8.00	8.01	8.08	8.06	8.10	7.96			
³ B ₁ (n → 3d)	8.48	8.47	8.48	8.48	8.41	8.40	8.42				
¹ A''[F](n → π*)	2.87	2.84	2.86	2.86	2.84	2.82	2.80				

state	methanimine								lit.	
	<i>aug-cc-pVDZ</i>				<i>aug-cc-pVTZ</i>				th. ^d	th. ^e
	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI			
¹ A''(n → π*)	5.26	5.24	5.25	5.25	5.20	5.19	5.23	5.32	5.18	
³ A''(n → π*)	4.63	4.63	4.63	4.63	4.61	4.61	4.65			

state	thioformaldehyde								lit.	
	<i>aug-cc-pVDZ</i>				<i>aug-cc-pVTZ</i>				exp. ^a	exp. ^f
	CC3	CCSDT	CCSDTQ	exFCI	CC3	CCSDT	exFCI			
¹ A ₂ (n → π*)	2.27	2.25	2.26	2.26	2.23	2.21	2.22		2.03	
¹ B ₂ (n → 4s)	5.80	5.80	5.82	5.83	5.91	5.89	5.96	5.85	5.84	
¹ A ₁ (π → π*)	6.62	6.60	6.51	6.5 ^g	6.48	6.47	6.4 ^g	6.2	5.54	
³ A ₂ (n → π*)	1.97	1.96	1.96	1.97	1.94	1.93	1.94		1.80	
³ A ₁ (π → π*)	3.43	3.43	3.44	3.45	3.38	3.38	3.43	3.28		
³ B ₂ (n → 4s)	5.64	5.63	5.65	5.66	5.72	5.71	5.6 ^g			
¹ A ₂ [F](n → π*)	2.00	2.00	1.98	1.98	1.97	1.98	1.95			

^aVarious experimental sources, summarized in ref 116. ^bMR-AQCC-LRT calculations from ref 134. ^cCC3/*aug-cc-pVQZ* calculations from ref 30. ^dDMC results from ref 135. ^eCCSDT/*aug-cc-pVTZ* calculations from ref 59. ^f0–0 energies collected in ref 136. ^gCI convergence too slow to provide reliable estimates. ^hAll values are in eV.

(12.78 eV) is probably more accurate than the one of ref 119 (12.90 eV), whereas the opposite is likely true for the highest ¹Π_u state that was reported to be located at 13.10 and 13.24 eV in refs 13 and 119, respectively. One could argue that reaching agreement between CI and CC is particularly challenging for these two states. However, performing the basis set extrapolation starting from the CCSDTQP/*aug-cc-pVDZ* results would yield similar TBE of 12.77 and 13.22 eV.

For the isoelectronic carbon monoxide, experimental vertical energies deduced from rovibronic data¹¹⁸ using a numerical approach are also available.^{22,120} With the *aug-cc-pVTZ* (*aug-cc-pVQZ*) atomic basis set, the CCSDT and CC3 results are within 0.02 eV (0.03 eV) and 0.03 eV (0.03 eV) of the exFCI results, whereas the errors made by both CCSDTQ and CCSDTQP are again trifling. As for dinitrogen, all the valence states are rather close from the basis set limit with *aug-cc-pVTZ*, whereas larger basis sets are required for the Rydberg states (Table S2). By correcting the exFCI/*aug-cc-pVQZ* (exFCI/*aug-cc-pVTZ* for the highest triplet state) data with basis set effects determined at the CC3/*d-aug-cc-pVSZ* level, we obtain TBE values that can be compared to the

experimental estimates. The computed MAD is 0.05 eV, the largest deviations being obtained for the Δ and Σ⁻ excited states of both spin symmetries. The agreement between theory and experiment is therefore very satisfying though slightly less impressive than for N₂. We note that the CC3/*aug-cc-pVTZ* C = O bond length (1.134 Å) is 0.006 Å larger than the experimental *r_e* value of 1.128 Å,¹¹⁸ whereas the discrepancy is twice smaller for dinitrogen: 1.101 Å for CC3/*aug-cc-pVTZ* compared to 1.098 Å experimentally. This might partially explained the larger deviations noticed for carbon monoxide.

3.3. Acetylene and Ethylene. Acetylene is the smallest conjugated organic molecule possessing stable low-lying excited-state structures, therefore allowing to investigate vertical fluorescence. This molecule has been the subject of previous investigations at the CASPT2,¹²⁴ CCSD,¹²⁵ CCSDT,⁵⁹ and MR-AQCC¹²⁶ levels. Our results are collected in Table 3. With the double-ζ basis set, the differences between the CC3, CCSDT, and CCSDTQ results are negligible, and the latter estimates are also systematically within 0.02 eV of the exFCI results. In contrast to water and ammonia, both CC3 and CCSDT provide similar accuracies compared to higher

Table 5. Vertical (Absorption) Transition Energies for Various Excited States of Diazomethane (Top) and Ketene (Bottom)^v

molecule	state	aug-cc-pVDZ			aug-cc-pVTZ			lit.	
		CC3	CCSDT	exFCI	CC3	CCSDT	exFCI	exp.	th.
acetaldehyde	¹ A'(n → π*)	4.34	4.32	4.34	4.31	4.29	4.31	4.27 ^a	4.29 ^b
	³ A'(n → π*)	3.96	3.95	3.98	3.95	3.94	4.0 ^c	3.97 ^a	3.97 ^b
cyclopropene	¹ B ₁ (σ → π*)	6.72	6.71	6.7 ^c	6.68	6.68	6.6 ^c	6.45 ^d	6.89 ^e
	¹ B ₂ (π → π*)	6.77	6.78	6.82	6.73	6.75	6.7 ^c	7.00 ^f	7.11 ^e
	³ B ₂ (π → π*)	4.34	4.35	4.35	4.34		4.38	4.16 ^f	4.28 ^g
	³ B ₁ (σ → π*)	6.43	6.43	6.43	6.40		6.45		6.40 ^g
diazomethane	¹ A ₂ (π → π*)	3.10	3.10	3.09	3.07	3.07	3.14	3.14 ^h	3.21 ⁱ
	¹ B ₁ (π → 3s)	5.32	5.35	5.35	5.45	5.48	5.54		5.33 ⁱ
	¹ A ₁ (π → π*)	5.80	5.82	5.79	5.84	5.86	5.90	5.9 ^h	5.85 ⁱ
	³ A ₂ (π → π*)	2.84	2.84	2.81	2.83	2.82	2.8 ^c		2.92 ^j
	³ A ₁ (π → π*)	4.05	4.04	4.03	4.03	4.02	4.05		3.97 ^j
	³ B ₁ (π → 3s)	5.17	5.20	5.18	5.31	5.34	5.35		
	³ A ₁ (π → 3p)	6.83	6.83	6.81	6.80		6.82		7.02 ^j
	¹ A' [F] (π → π*)	0.68	0.67	0.65	0.68	0.67	0.71		
formamide	¹ A'(n → π*)	5.71	5.68	5.70	5.66	5.63	5.7 ^c	5.8 ^k	5.63 ^l
	¹ A'(n → 3s)	6.65	6.64	6.67	6.74	6.74		6.35 ^k	6.62 ^l
	¹ A' (π → π*) ^m	7.63	7.62	7.64	7.62		7.63	7.37 ^k	7.22 ^l
	¹ A'(n → 3p) ^m	7.31	7.29		7.40	7.38		7.73 ^k	7.66 ^l
	³ A'(n → π*)	5.42	5.39	5.42	5.38		5.4 ^c	5.2 ^k	5.34 ^l
	³ A' (π → π*)	5.83	5.81	5.82	5.82		5.7 ^c	~6 ^k	5.74 ^l
ketene	¹ A ₂ (π → π*)	3.89	3.88	3.84	3.88	3.87	3.86	3.7 ⁿ	3.74 ^o
	¹ B ₁ (n → 3s)	5.83	5.86	5.88	5.96	5.99	6.01	5.86 ⁿ	5.82 ^o
	¹ A ₂ (π → 3p)	7.05	7.09	7.08	7.16	7.20	7.18		7.00 ^o
	³ A ₂ (n → π*)	3.79	3.78	3.79	3.78	3.78	3.77	3.8 ^p	3.62 ^q
	³ A ₁ (π → π*)	5.62	5.61	5.64	5.61	5.60	5.61	5 ^p	5.42 ^q
	³ B ₁ (n → 3s)	5.63	5.66	5.68	5.76	5.80	5.79	5.8 ^p	5.69 ^q
	³ A ₂ (π → 3p)	7.01	7.05	7.07	7.12	7.17	7.12		
	¹ A' [F] (π → π*)	1.00	0.99	0.96	1.00	1.00	1.00		
nitrosomethane	¹ A'(n → π*)	2.00	1.98	1.99	1.96	1.95	2.0 ^c	1.83 ^r	1.76 ^s
	¹ A' (n, n → π*, π*)	5.75	5.26	4.81	5.76	5.29	4.72		4.96 ^s
	¹ A' (n → 3s/3p)	6.20	6.19	6.29	6.31	6.30	6.4 ^c		6.54 ^s
	³ A'(n → π*)	1.13	1.12	1.15	1.14	1.13	1.16		1.42 ^t
	³ A' (π → π*)	5.54	5.54	5.56	5.51		5.60		5.55 ^t
	¹ A' [F] (n → π*)	1.70	1.69	1.70	1.69	1.66	1.7 ^c		
streptocyanine-C1	¹ B ₂ (π → π*)	7.14	7.12	7.14	7.13	7.11	7.1 ^c		7.16 ^u
	³ B ₂ (π → π*)	5.48	5.47	5.47	5.48	5.47	5.52		

^aElectron impact experiment from ref 145. ^bNEVPT-PC from ref 127. ^cCI convergence too slow to provide more reliable estimates. ^dMaximum in the gas UV from ref 146. ^eCCSDT/TZVP from ref 58. ^fElectron impact experiment from ref 147. ^gCC3/aug-cc-pVTZ from ref 32. ^hVUV maxima from ref 148. ⁱCCSD/6-311(3+,+)G(d) calculations from ref 149. ^jMR-CC/DZP calculations from ref 150. ^kEELS (singlet) and trapped electron (triplet) experiments from ref 151. ^lnR-SI-CCSD(T) results from ref 142. ^mStrong state mixing. ⁿElectron impact experiment from ref 152. ^oCASPT2/6-311+G(d) results from ref 153. ^pElectron impact experiment from ref 116. ^qSTEOM-CCSD/Sad+//CCSD/Sad+ results from ref 154. ^rMaximum in the gas UV from ref 155. ^sCASPT2/ANO results from ref 156. ^tCASSCF/cc-pVDZ results from ref 157. ^uexCC3//MP2 result from ref 128. ^vAll values are in eV.

levels of theory. As expected, for valence states, going from double- to triple- ζ basis set tends to slightly decrease the computed energies (except for the lowest triplet). Nonetheless, as with the smaller basis set, the same near-perfect methodological match pertains with aug-cc-pVTZ. Estimating the exFCI/aug-cc-pVTZ results from the exFCI/aug-cc-pVDZ values and CC3 basis set effects would yield estimates with absolute errors of 0.00–0.02 eV. One also notices that the exFCI/aug-cc-pVTZ values are all extremely close to the previous MR-AQCC estimates, whereas the published CASPT2 values appear to be too low though closer from the electron impact experiment, underlying once more the difficulty to obtain very accurate experimental estimates for

vertical energies. This underestimating trend of standard CASPT2 was reported before for other molecules.^{127,128} Although our theoretical vertical energy estimates still slightly vary when passing from the aug-cc-pVDZ to aug-cc-pVTZ basis sets, we claim that these vertical energies are probably more trustworthy for further benchmarks than the available experimental values because basis set effects beyond aug-cc-pVTZ seem rather limited (Table S3).

Despite its small size, ethylene remains a challenging molecule and is included in many benchmark sets.^{17,26,29,30,75,132} The assignments of the experimental data have been the subject of countless works, and we refer the interested readers to the discussions in refs 30, 91, 116, 130,

131, and 133. On the theoretical side, the most complete and accurate investigation dedicated to the excited states of ethylene is due to Davidson's group, who performed refined CI calculations.¹³¹ They indeed obtained highly accurate transition energies for ethylene, including for the valence yet challenging $^1B_{1u}$ state. From our data, collected in Table 3, one notices that the differences between exFCI/*aug-cc-pVDZ* and CCSDTQ/*aug-cc-pVDZ* results are again trifling, the largest deviation being obtained for the $^3B_{3u}(\pi \rightarrow 3s)$ Rydberg state (0.02 eV). In addition, given the nice agreement between CC3, CCSDT, and exFCI values, one can directly compare our CC3/*aug-cc-pVSZ* results (Table S3) to the values of reported in ref 131: a mean absolute deviation (MAD) of 0.03 eV is obtained. The fact that our transition energies tend to be slightly smaller than Davidson's is likely due to geometrical effects. Indeed, our CC3/*aug-cc-pVTZ* C=C distance is 1.3338 Å, i.e., slightly longer than the best estimate provided in Davidson's work (1.3305 Å). Recently, a stochastic heat-bath CI (SHCI)/ANO-L-pVTZ work reported 4.59 and 8.05 eV values for the $^3B_{1u}$ and $^1B_{1u}$ states, respectively,⁹¹ and we also ascribe the differences with our results to the use of a MP2 geometry in ref 91. Interestingly, these authors found quite large discrepancies between their SHCI and their CC results. Indeed, they reported CR-EOMCC(2,3)D estimates significantly larger than their SHCI results with +0.17 and +0.20 eV upshifts for the triplet and singlet states, respectively. This highlights that only high-level CC schemes are able to recover the exFCI (or SHCI) results for ethylene.

3.4. Formaldehyde, Methanimine, and Thioformaldehyde. Similarly to ethylene, formaldehyde is a very popular test molecule,^{17,22,26,29,30,59,75,76,132,137–142} and stands as the prototype carbonyl dye with a low-lying $n \rightarrow \pi^*$ transition. Nevertheless, even for this particular valence state, well-separated from higher-lying excited states, the choice of an experimental reference remains difficult. Indeed, values of 3.94,²² 4.00,^{26,29,138} 4.07,^{17,75,139} and 4.1 eV,^{137,140} have been used in previous theoretical benchmarks. In contrast to their oxygen cousin, both methanimine and thioformaldehyde were the subject of less attention from the theoretical community.^{135,143,144} The results obtained for these three molecules are collected in Table 4. Considering all transitions listed in this table, one obtains a MAD of 0.01 eV between the CCSDTQ/*aug-cc-pVDZ* and exFCI/*aug-cc-pVDZ* results, the largest discrepancies of 0.03 eV being observed for two states for which the differences between CCSDT and CCSDTQ are also large (0.05 eV). As in water, using the exFCI/*aug-cc-pVDZ* values as reference, we found that CC3 delivers slightly more accurate transition energies (MAD of 0.02 eV, maximal deviation of 0.06 eV) than CCSDT (MAD of 0.03 eV, maximal deviation of 0.07 eV). By adding the difference between CC3/*aug-cc-pVTZ* and CC3/*aug-cc-pVDZ* results to the exFCI/*aug-cc-pVDZ* values, we obtain good estimates of the actual exFCI/*aug-cc-pVTZ* data, with a MAD of 0.02 eV for formaldehyde. Compared to the CC3/*aug-cc-pVQZ* results of Thiel,³⁰ the transition energies reported in Table 4 are slightly larger, which is probably due to the influence of the ground-state geometry rather than to basis set effects (see Table S4). Indeed, the carbonyl bond is significantly more contracted with CC3/*aug-cc-pVTZ* (1.208 Å) than with MP2/6-31G(d) (1.221 Å). In particular, for the hallmark $n \rightarrow \pi^*$, our best estimate is 3.97 eV (*vide infra*), nicely matching a previous MR-AQCC value of 3.98 eV,¹³⁴ but significantly below the previous DMC/BLYP estimate of 4.24 eV.¹³⁵ The

latter discrepancy is probably due to the use of both different structures and pseudopotentials within DMC calculations.

For methanimine and thioformaldehyde, the basis set effects are rather small for the states considered here (see Table S4) and the data reported in the present work are probably the most accurate vertical transition energies reported to date. For the latter molecule, these vertical estimates are systematically larger than the known experimental 0–0 energies,¹³⁶ which is the expected trend.

3.5. Larger Compounds. Let us now turn our attention to molecules that encompass three heavy (non-hydrogen) atoms. We have treated seven molecules of that family, and all were previously investigated at several levels of theory: acetaldehyde,^{26,29,127,138–140,158,159} cyclopropene,^{30–32,58,132,160} diazomethane,^{149,150,158,161} formamide,^{30–32,58,59,162,163} ketene,^{150,153,154,164} nitrosomethane,^{156,157,165,166} and the shortest streptocyanine.^{128,167–170} The results are gathered in Table 5. Note that, for these molecules containing three heavy atoms, it is sometimes challenging to obtain reliable exFCI estimates, especially for the largest basis set.

Experimentally, the lowest singlet and triplet $n \rightarrow \pi^*$ transitions of acetaldehyde are located 0.3–0.4 eV above their formaldehyde counterparts,^{116,145} and this trend is accurately reproduced by theory, which also delivers estimates very close to the NEVPT2 values given in ref 127.

For cyclopropene, the lowest singlet $\sigma \rightarrow \pi^*$ and $\pi \rightarrow \pi^*$ are close from one another, and both CCSDT and exFCI predict the former to be slightly more stabilized, which is consistent with the large basis set CC3 results obtained by Thiel.³²

For the isoelectronic diazomethane and ketene molecules (see Table 5), one notes, yet again, consistent results with, however, differences between the exFCI/*aug-cc-pVTZ* and CCSDT/*aug-cc-pVTZ* results larger than 0.05 eV for the two lowest singlet states of diazomethane. There is also a reasonable match between our data and previous theoretical results reported for these two molecules.^{149,150,153,154} The basis set effects are significant for the Rydberg transitions, especially for the $\pi \rightarrow 3s$ states of diazomethane (Table S5).

In formamide, we found strong state mixing between the lowest singlet valence and Rydberg states of A' symmetry. This is consistent with the CCSDT/TZVP analysis of Kannar and Szalay,⁵⁸ who reported, for example, a larger oscillator strength for the lowest Rydberg state than for the $\pi \rightarrow \pi^*$ transition. This state-mixing problem pertains with *aug-cc-pVTZ*, making unambiguous assignments difficult. Consequently, we have decided to classify the three lowest $^1A'$ transitions according to their dominant orbital character, which gives a picture consistent with the computed oscillator strengths (*vide infra*) but yields state inversions compared to Thiel's and Szalay's assignments.^{31,58} This strong state mixing also prevented the convergence of several state energies with the exFCI/*aug-cc-pVTZ* approach. Despite these uncertainties, we obtained transition energies for the Rydberg states that are much closer from experiment¹⁵¹ as well as from previous multireference CC estimates,¹⁴² than the TZVP ones.⁵⁸

Nitrosomethane is an interesting test molecule for three reasons: (i) it presents very low-lying $n \rightarrow \pi^*$ states of A'' symmetry, close to ca. 2.0 eV (singlet) and 1.2 eV (triplet), among the smallest absorption energies found in a compact molecule;¹⁷¹ (ii) it changes from an eclipsed to a staggered conformation of the methyl group when going from the ground to the lowest singlet state;^{157,172,173} (iii) the lowest-lying singlet A' state corresponds to an almost pure double

Table 6. TBE (in eV) for Various States and Wave Function Approaches^e

	state	<i>f</i>	% <i>T</i> ₁	TBE(FC) AVTZ	corrected TBE		
					method	corr.	value
acetaldehyde	¹ A''(V; n → π*)	0.000	91.3	4.31	exFCI/AVTZ	AVQZ	4.31
	³ A''(V; n → π*)		97.9	3.97 ^a	exFCI/AVDZ	AVQZ	3.98
acetylene	¹ Σ _g ⁻ (V; π → π*)		96.5	7.10	exFCI/AVTZ	dAVSZ	7.10
	¹ Δ _g (V; π → π*)		93.3	7.44			7.44
	³ Σ _g ⁺ (V; π → π*)		99.2	5.53			5.56
	³ Δ _g (V; π → π*)		99.0	6.40			6.40
	³ Σ _g ⁻ (V; π → π*)		98.8	7.08			7.09
	¹ A _u [F](V; π → π*)		95.6	3.64			3.63
	¹ A ₂ [F](V; π → π*)		95.5	3.85			3.85
ammonia	¹ A ₂ (R; n → 3s)	0.086	93.5	6.59	exFCI/AVQZ	dAVSZ	6.66
	¹ E(R; n → 3p)	0.006	93.7	8.16			8.21
	¹ A ₁ (R; n → 3p)	0.003	94.0	9.33			8.65
	¹ A ₂ (R; n → 4s)	0.008	93.6	9.96	exFCI/AVTZ	dAVSZ	9.19
carbon monoxide	³ A ₂ (R; n → 3s)		98.2	6.31	exFCI/AVQZ	dAVSZ	6.37
	¹ Π(V; n → π*)	0.084	93.1	8.49	exFCI/AVQZ	dAVSZ	8.48
	¹ Σ ⁻ (V; π → π*)		93.3	9.92			9.98
	¹ Δ(V; π → π*)		91.8	10.06			10.10
	¹ Σ ⁺ (R)	0.003	91.5	10.95			10.80
	¹ Σ ⁺ (R)	0.200	92.9	11.52			11.42
	¹ Π(R)	0.053	92.4	11.72			11.55
	³ Π(V; n → π*)		98.7	6.28			6.28
	³ Σ ⁺ (V; π → π*)		98.7	8.45			8.49
	³ Δ(V; π → π*)		98.4	9.27			9.28
	³ Σ ⁻ (V; π → π*)		97.5	9.80			9.77
	³ Σ ⁺ (R)		98.0	10.47	exFCI/AVTZ	dAVSZ	10.37
	cyclopropene	¹ B ₁ (V; σ → π*)	0.001	92.8	6.68 ^b	CCSDT/AVTZ	AVQZ
¹ B ₂ (V; π → π*)		0.071	95.1	6.79 ^a	exFCI/AVDZ	AVQZ	6.78
³ B ₂ (V; π → π*)			98.0	4.38	exFCI/AVTZ	AVQZ	4.38
³ B ₁ (V; σ → π*)			98.9	6.45			6.45
diazomethane	¹ A ₂ (V; π → π*)		90.1	3.14	exFCI/AVTZ	dAVQZ	3.13
	¹ B ₁ (R; π → 3s)	0.002	93.8	5.54			5.59
	¹ A ₁ (V; π → π*)	0.210	91.4	5.90			5.89
	³ A ₂ (V; π → π*)		97.7	2.79 ^a	exFCI/AVDZ	dAVQZ	2.80
	³ A ₁ (V; π → π*)		98.6	4.05	exFCI/AVTZ	dAVQZ	4.05
	³ B ₁ (R; π → 3s)		98.0	5.35			5.40
	³ A ₁ (R; π → 3p)		98.5	6.82			6.72
	¹ A'' [F](V; π → π*)		87.4	0.71			0.70
dinitrogen	¹ Π _g (V; n → π*)		92.6	9.34	exFCI/AVQZ	dAVSZ	9.33
	¹ Σ _g ⁻ (V; π → π*)		97.2	9.88			9.91
	¹ Δ _g (V; π → π*)	0.000	95.9	10.29			10.31
	¹ Σ _g ⁺ (R)		92.2	12.98			12.30
	¹ Π _u (R)	0.229	82.9	13.03	exFCI/AVTZ	dAVSZ	12.73
	¹ Σ _u ⁺ (R)	0.296	92.8	13.09			12.95
	¹ Π _u (R)	0.000	87.4	13.46			13.27
	³ Σ _u ⁺ (V; π → π*)		99.3	7.70	exFCI/AVQZ	dAVSZ	7.74
	³ Π _g (V; n → π*)		98.4	8.01			8.03
	³ Δ _g (V; π → π*)		99.3	8.87			8.88
	³ Σ _u ⁻ (V; π → π*)		98.8	9.66			9.65
ethylene	¹ B _{3u} (R; π → 3s)	0.078	95.1	7.39	exFCI/AVTZ	dAVSZ	7.43
	¹ B _{1u} (V; π → π*)	0.346	95.8	7.93			7.92
	¹ B _{1g} (R; π → 3p)		95.3	8.08			8.10
	³ B _{1u} (V; π → π*)		99.1	4.54			4.54
	³ B _{3u} (R; π → 3s)		98.5	7.23 ^a	exFCI/AVDZ	dAVSZ	7.28
	³ B _{1g} (R; π → 3p)		98.4	7.98 ^a			8.00
	¹ A ₂ (V; n → π*)		91.5	3.98	exFCI/AVTZ	dAVSZ	3.97
formaldehyde	¹ B ₂ (R; n → 3s)	0.021	91.7	7.23			7.30

Table 6. continued

	state	<i>f</i>	% <i>T</i> ₁	TBE(FC) AVTZ	corrected TBE		
					method	corr.	value
	¹ B ₂ (R; n → 3p)	0.037	92.4	8.13			8.14
	¹ A ₁ (R; n → 3p)	0.052	91.9	8.23			8.27
	¹ A ₂ (R; n → 3p)		91.7	8.67			8.50
	¹ B ₁ (V; σ → π*)	0.001	90.8	9.22			9.21
	¹ A ₁ (V; π → π*)	0.135	90.4	9.43			9.26
	³ A ₂ (V; n → π*)		98.1	3.58			3.58
	³ A ₁ (V; π → π*)		99.0	6.06			6.07
	³ B ₂ (R; n → 3s)		97.1	7.06			7.14
	³ B ₂ (R; n → 3p)		97.4	7.94			7.96
	³ A ₁ (R; n → 3p)		97.2	8.10			8.15
	³ B ₁ (R; n → 3d)		97.9	8.42			8.42
	¹ A'' [F] (V; n → π*)		87.8	2.80			2.80
formamide	¹ A'' (V; n → π*)	0.000	90.8	5.65 ^a	exFCI/AVDZ	AVQZ	5.63
	¹ A' (R; n → 3s)	0.001	88.6	6.77 ^a			6.81
	¹ A' (V; π → π*)	0.251	89.3	7.63	exFCI/AVTZ	AVQZ	7.64
	¹ A' (R; n → 3p)	0.111	89.6	7.38 ^b	CCSDT/AVTZ	AVQZ	7.41
	³ A'' (V; n → π*)		97.7	5.38 ^c	exFCI/AVDZ	AVQZ	5.37
	³ A' (V; π → π*)		98.2	5.81 ^c			5.81
hydrogen chloride	¹ Π(CT)	0.056	94.3	7.84	exFCI/AVQZ	dAVSZ	7.86
hydrogen sulfide	¹ A ₂ (R; n → 4p)		94.6	6.18	exFCI/AVQZ	dAVSZ	6.10
	¹ B ₁ (R; n → 4s)	0.063	94.3	6.24			6.29
	³ A ₂ (R; n → 4p)		98.7	5.81			5.74
	³ B ₁ (R; n → 4s)		98.4	5.88			5.94
ketene	¹ A ₂ (V; π → π*)		91.0	3.86	exFCI/AVTZ	dAVQZ	3.86
	¹ B ₁ (R; n → 3s)	0.035	93.9	6.01			6.06
	¹ A ₂ (R; π → 3p)		94.4	7.18			7.19
	³ A ₂ (V; n → π*)		91.0	3.77			3.77
	³ A ₁ (V; π → π*)		98.6	5.61			5.60
	³ B ₁ (R; n → 3s)		98.1	5.79			5.85
	³ A ₂ (R; π → 3p)		94.4	7.12			7.14
	¹ A'' [F] (V; π → π*)		87.9	1.00			1.00
methanimine	¹ A'' (V; n → π*)	0.003	90.7	5.23	exFCI/AVTZ	dAVQZ	5.21
	³ A'' (V; n → π*)		98.1	4.65			4.64
nitrosomethane	¹ A'' (V; n → π*)	0.000	93.0	1.96 ^d	exFCI/AVDZ	AVQZ	1.95
	¹ A' (V; n, n → π*, π*)	0.000	2.5	4.72	exFCI/AVTZ	AVQZ	4.69
	¹ A' (R; n → 3s/3p)	0.006	90.8	6.40 ^d	exFCI/AVDZ	AVQZ	6.42
	³ A'' (V; n → π*)		98.4	1.16			1.16
	³ A' (V; π → π*)		98.9	5.60			5.61
	¹ A'' [F] (V; n → π*)		92.7	1.67 ^a	exFCI/AVDZ	AVQZ	1.66
streptocyanine-C1	¹ B ₂ (V; π → π*)	0.347	88.7	7.13 ^a	exFCI/AVDZ	AVQZ	7.12
	³ B ₂ (V; π → π*)		98.3	5.52	exFCI/AVTZ	AVQZ	5.52
thioformaldehyde	¹ A ₂ (V; n → π*)		89.3	2.22	exFCI/AVTZ	dAVQZ	2.20
	¹ B ₂ (R; n → 4s)	0.012	92.3	5.96			5.99
	¹ A ₁ (V; π → π*)	0.178	90.8	6.38 ^d	CCSDTQ/AVDZ	dAVQZ	6.34
	³ A ₂ (V; n → π*)		97.7	1.94	exFCI/AVTZ	dAVQZ	1.94
	³ A ₁ (V; π → π*)		98.9	3.43			3.44
	³ B ₂ (R; n → 4s)		97.6	5.72 ^a	exFCI/AVDZ	dAVQZ	5.76
	¹ A ₂ [F] (V; n → π*)		87.2	1.95	exFCI/AVTZ	dAVQZ	1.94
water	¹ B ₁ (R; n → 3s)	0.054	93.4	7.62	exFCI/AVQZ	dAVSZ	7.70
	¹ A ₂ (R; n → 3p)		93.6	9.41			9.47
	¹ A ₁ (R; n → 3s)	0.100	93.6	9.99			9.97
	³ B ₁ (R; n → 3s)		98.1	7.25			7.33
	³ A ₂ (R; n → 3p)		98.0	9.24			9.30
	³ A ₁ (R; n → 3s)		98.2	9.54			9.59

^aexCI/aug-cc-pVDZ data corrected with the difference between CCSDT/aug-cc-pVTZ and CCSDT/aug-cc-pVDZ values. ^bCCSDT/aug-cc-pVTZ value. ^cexCI/aug-cc-pVDZ data corrected with the difference between CC3/aug-cc-pVTZ and CC3/aug-cc-pVDZ values. ^dCCSDTQ/aug-cc-

Table 6. continued

pVDZ data corrected with the difference between CCSDT/*aug-cc-pVTZ* and CCSDT/*aug-cc-pVDZ* values. ^eFor each state, we provide the oscillator strength and percentage of single excitations obtained at the CC3(FC)/*aug-cc-pVTZ* level. Unless otherwise stated, the TBE(FC)/*aug-cc-pVTZ* have been obtained directly from exFCI. For the basis-set-corrected TBE, we provide the method used to determine the starting value and the basis set used at the CC3(full) level to correct it. CC3(full)/*aug-cc-pVTZ* geometries and abbreviated forms of Dunning's basis set are systematically used. R, V and F stand for Rydberg, valence and fluorescence, respectively.

excitation of $(n, n) \rightarrow (\pi^*, \pi^*)$ nature.¹⁵⁶ Indeed, CC3 returns a 2.5% single excitation character only for this second transition, to be compared to more than 80% (and generally more than 90%) in all other states treated in this work (*vide infra*). For example, the notoriously difficult A_g dark state of butadiene has a 72.8% single character.³⁰ For the A'' state of nitrosomethane, CC3, CCSDT and exFCI yield similar results, and the corresponding transition energies are slightly larger than previous CASPT2 estimates.¹⁵⁶ In contrast, the CC approaches are expectedly far from the spot for the $(n, n) \rightarrow (\pi^*, \pi^*)$ transition: they yield values significantly blue-shifted and large discrepancies between the CC3 and CCSDT values are found. For this particular state, it is not surprising that the exFCI result is indeed closer to the CASPT2 value,¹⁵⁶ as modeling double excitations with single-reference CC models is not a natural choice.

Finally for the shortest model cyanine, a molecule known to be difficult to treat with TD-DFT,¹⁷⁰ all the theoretical results given in Table 5 closely match each other for both the singlet and triplet manifolds. For the former, the reported CASPT2 (with IPEA) value of 7.14 eV also fits these estimates.¹²⁸

3.6. Theoretical Best Estimates. We now turn to the definition of theoretical best estimates. We decided to provide two sets for these estimates, one obtained in the frozen-core approximation with the *aug-cc-pVTZ* atomic basis set, and one including further corrections for basis set and “all electron” (full) effects. This choice allows further benchmarks to either consider a reasonably compact basis set, therefore allowing to test many levels of theory, or to rely on values closer to the basis set limit. For the former set, we systematically selected exFCI/*aug-cc-pVTZ* values except when explicitly stated. For the latter set, both the “all electron” correlation and the basis set corrections (see Supporting Information for complete data) were systematically obtained at the CC3 level of theory and used *d-aug-cc-pV5Z* for the nine smallest molecules, but slightly more compact basis sets for the larger compounds. At least for Rydberg states, the use of *d-aug-cc-pVQZ* apparently delivers results closer to basis set convergence than *aug-cc-pV5Z*, and the former basis set was used when technically possible. The interested readers may find in Supporting Information the values obtained with and without applying the frozen-core approximation for several basis sets. Clearly, the largest amount of the total correction originates from basis set effects. In other words, “full” and frozen-core transition energies are typically within 0.01–0.02 eV of each other for a given basis set. The results are listed in Table 6 and provide a total of 110 transition energies. This set of states is rather diverse with 61 singlet and 45 triplet states, 60 valence and 45 Rydberg states, 21 $n \rightarrow \pi^*$ and 38 $\pi \rightarrow \pi^*$ states, with an energetic span from 0.70 to 13.27 eV. Among these 110 excitation energies, only 13 are characterized by a single-excitation character smaller than 90% according to CC3. As expected,³⁰ the dominant single-excitation character is particularly pronounced for triplet excited states. Therefore, this set is adequate for evaluating single-reference methods, though a few challenging cases are incorporated. Conse-

quently, we think that the TBE listed in Table 6 contribute to fulfill the need of more accurate reference excited state energies, as pointed out by Thiel one decade ago.³⁰ However, the focus on small compounds and the lack of charge-transfer states constitute significant biases in the present set of transition energies.

3.7. Benchmarks. We have used the TBE(FC)/*aug-cc-pVTZ* benchmark values to assess the performances of 12 wave function approaches, namely, ADC(2), ADC(3), CIS(D), CIS(D_∞), CC2, STEOM-CCSD, CCSD, CCSDR(3), CCSDT-3, CC3, CCSDT, and CCSDTQ. The complete list of results can be found in Table S6 in the Supporting Information. As expected, only the approaches including iterative triples, that is, ADC(3), CCSDT-3, CC3, and CCSDT are able to predict the presence of the doubly excited $(n, n) \rightarrow (\pi^*, \pi^*)$ transition in nitrosomethane (see Tables 5 and S6), but they all yield large quantitative errors. Indeed, the TBE value of 4.72 eV is strongly underestimated by ADC(3) (3.00 eV) and significantly overshoot by the three CC models with estimates of 6.02 eV, 5.76, and 5.29 eV with CCSDT-3, CC3, and CCSDT, respectively. This 0.26 eV difference between the CCSDT-3 and CC3 values is also the largest discrepancy between these two models in the tested set. Obviously, from a general perspective, one should not use the standard single-reference wave function methods to describe double excitations. Therefore, the $(n, n) \rightarrow (\pi^*, \pi^*)$ transition of nitrosomethane was removed from our statistical analysis. Likewise, for the three lowest $^1A'$ excited states of formamide, strong state mixing — involving two or three states — are found at all levels of theory, making unambiguous assignments impossible. Consequently, they are also excluded from our statistics.

In Table 7, we report, for the entire set of compounds, the mean signed error (MSE), mean absolute error (MAE) root-mean-square deviation (RMS), as well as the positive [Max(+)] and negative [Max(-)] maximum deviations. A graphical representation of the errors obtained with all methods can be found in Figure 1. Note that only singlet states could be computed with the programs used for CCSDR(3) and CCSDT-3. As shown in Figure 1, CCSDTQ is on the spot with tiny MSE and MAE, which is consistent with the analysis carried out for individual molecules. With this method, the negative and positive maximum deviations are as small as -0.05 eV (singlet $n \rightarrow 4s$ Rydberg transition of thioformaldehyde) and +0.06 eV ($^1\Sigma_u^+$ Rydberg transition of dinitrogen), respectively. The three other CC models with iterative triples (CCSDT-3, CC3, and CCSDT) also deliver extremely accurate transition energies with MAE of 0.03 eV only. In agreement with the analysis of Watson and co-workers, we do not find any significant (statistical) differences between CCSDT-3 and CC3,⁵⁷ and although the former theory is formally closer to CCSDT, it does not seem more advantageous nor disadvantageous than CC3 in practice. The very good performance of CC3 is also consistent with the analysis of Thiel and co-workers, who reported a strong agreement with CASPT2,³² as well as with the conclusion of Szalay's group

Table 7. Mean Signed Error (MSE), Mean Absolute Error (MAE), Root-Mean Square Deviation (RMS), Positive [Max(+)] and Negative [Max(-)] Maximal Deviations with Respect to TBE(FC)/*aug-cc-pVTZ* for the Transition Energies Listed in Table S6^a

method	no. of states	MSE	MAE	RMS	Max(+)	Max(-)
CIS(D)	106	0.10	0.25	0.32	-0.63	1.06
CIS(D _∞)	106	-0.01	0.21	0.28	-0.76	0.57
CC2	106	0.03	0.22	0.28	-0.71	0.63
STEOM-CCSD	102	0.01	0.10	0.14	-0.56	0.40
CCSD	106	0.05	0.08	0.11	-0.17	0.40
CCSDR(3)	59	0.01	0.04	0.05	-0.07	0.25
CCSDT-3	58	0.01	0.03	0.05	-0.07	0.24
CC3	106	-0.01	0.03	0.04	-0.09	0.19
CCSDT	104	-0.01	0.03	0.03	-0.10	0.11
CCSDTQ	73	0.00	0.01	0.02	-0.05	0.06
ADC(2)	106	-0.01	0.21	0.28	-0.76	0.57
ADC(3)	106	-0.15	0.23	0.28	-0.79	0.39

^aAll values are in eV and have been obtained with the *aug-cc-pVTZ* basis set.

who found it very close to CCSDT.⁵⁹ Nevertheless, CCSDT is not, on average, significantly more accurate than CC3 nor CCSDT-3. In other words, CCSDT is probably not a sufficiently accurate benchmark to estimate the accuracy of CCSDT-3 nor CC3. The perturbative inclusion of triples via CCSDR(3) stands as a good compromise between computational cost and accuracy with a MAE of 0.04 eV, a conclusion also drawn in the benchmark study performed by Sauer and co-workers.⁵¹ These very small average deviations are related to the fact that the majority of our set is constituted of large single-excitation character transitions (see % T_1 in Table 6). Reasonably, we predict that they would slightly deteriorate for larger compounds.

For the second-order CC series, as expected, the errors increase when one uses more approximate models. Indeed, the MAE are 0.08, 0.10, and 0.22 eV with CCSD, STEOM-CCSD, and CC2, respectively. The magnitude of the CC2 average deviation is consistent with previous estimates obtained for Thiel's set (0.29 eV for singlets and 0.18 eV for triplets),³⁰ for fluorescence energies (0.21 eV for 12 small compounds),¹⁷⁴ as well as for larger compounds (0.15 eV for 0–0 energies of conjugated dyes).¹¹ Likewise, the fact that CCSD tends to overestimate the transition energies (positive MSE) was also reported previously in several works.^{26,30,57–59,159,174} It can be seen that Nooijen's STEOM approach, which was much less benchmarked previously, delivers an accuracy comparable to CCSD, with a smaller MSE but a large dispersion. More surprisingly, we found a MAE smaller with CCSD than with CC2, which contrasts with the results reported for Thiel's set,⁵¹ but is consistent with Kannar, Tajti and Szalay conclusion.⁵⁹ We attribute this effect to the small size of the compounds treated herein. Indeed, analyzing the TZVP values of ref 30., it appears clearly that CC2 more regularly outperforms CCSD for larger compounds.

As expected, the results for CIS(D_∞) and ADC(2), two closely related theories,^{6,144} are nearly equivalent, with only 4 (out of 106) cases for which a difference of 0.01 eV could be evidenced (Table S6). In addition, Table 7 evidences that ADC(2) provides an accuracy similar to CC2 for a smaller computational cost, whereas CIS(D) is slightly less accurate.

Both outcomes perfectly fit previous benchmarks.^{10,11,48,144,174} Conversely, we found that ADC(3) results are rather poor with average deviations larger than the ones obtained with ADC(2) and a clear tendency to provide red-shifted transition energies with a MSE of -0.15 eV. This observation is in sharp contrast with a previous investigation which concluded that ADC(3) and CC3 have very similar performances,⁴⁸ though the ADC(3) excitation energies were also found to be, on average, smaller by 0.20 eV compared to their CC3 counterparts. At this stage, it is difficult to know if the large MAE of ADC(3) reported in Table 7 originates solely from the small size of the compounds treated herein. However, the fact that the CCSD MSE is relatively small compared to previous benchmarks hints that the choice of compact compounds has a non-negligible effect on the statistics.

Let us analyze the ADC(3) errors more thoroughly. First, ADC(3) deviations are quite large for all subsets (*vide infra*). Second, we have found that, for the 46 transition energies for which ADC(2) yields an absolute error exceeding 0.15 eV compared to our TBE, the signs of the ADC(2) and ADC(3) errors systematically differ (see Figure 2); i.e., ADC(3) goes in the right "direction" but has the tendency to overcorrect ADC(2). This is clearly reminiscent of the well-known oscillating behavior of the Møller–Plesset perturbative series for ground state properties. Third, this overestimation of the corrections pertains for the states in which the ADC(2) absolute error is smaller than 0.15 eV. Indeed, in those 60 cases, there are only 10 transitions for which the ADC(3) values are more accurate than their second-order counterpart. As a consequence, taking the average between the ADC(2) and ADC(3) transition energies yield rather accurate estimates with a MAE as small as 0.10 eV for the full set, half of the MAE obtained with the parent methods.

We provide a more detailed analysis for several subsets of states in Table S7 in the Supporting Information. Globally, we found no significant difference between the singlet and triplet transitions, though all CC models (except STEOM-CCSD) provide slightly smaller deviations for the latter transitions, in line with their larger single-excitation character. With the computationally lighter methods, CIS(D), CIS(D_∞), ADC(2), and CC2, the MAEs are significantly smaller for the valence transitions (0.20, 0.15, 0.15, and 0.18 eV, respectively) than for the Rydberg transitions (0.32, 0.29, 0.29, and 0.26 eV, respectively). We also found MSE of opposite sign for valence and Rydberg transitions with CC2, which fits the results of Kannar and co-workers.⁵⁹ Surprisingly, ADC(3) gives 0.28 and 0.17 eV MAE for valence and Rydberg, respectively. All CC methods including triples deliver similar deviations for both sets of states. All methods provide smaller (or equal) MAE for the $n \rightarrow \pi^*$ than for the $\pi \rightarrow \pi^*$ transitions, which was already found for Thiel's set.³⁰ The differences are particularly significant with CIS(D), CC2, STEOM-CCSD, and ADC(3) with errors twice larger for $\pi \rightarrow \pi^*$ than $n \rightarrow \pi^*$ states. Finally, when considering the few states with % T_1 smaller than 90%, we logically found larger statistical errors with, for example, MAE of, e.g., 0.03 eV for CCSDTQ, 0.04 eV for CC3, and 0.06 eV for CCSDT-3.

3.8. On the Use of a Compact Basis Set. In several of the molecules considered here, we have found that adding corrections for basis set effects determined at the CC3 level to exFCI/*aug-cc-pVDZ* results effectively provides accurate estimates of the exFCI values directly determined with larger bases. Nevertheless, the dreadful scalings of both exFCI and

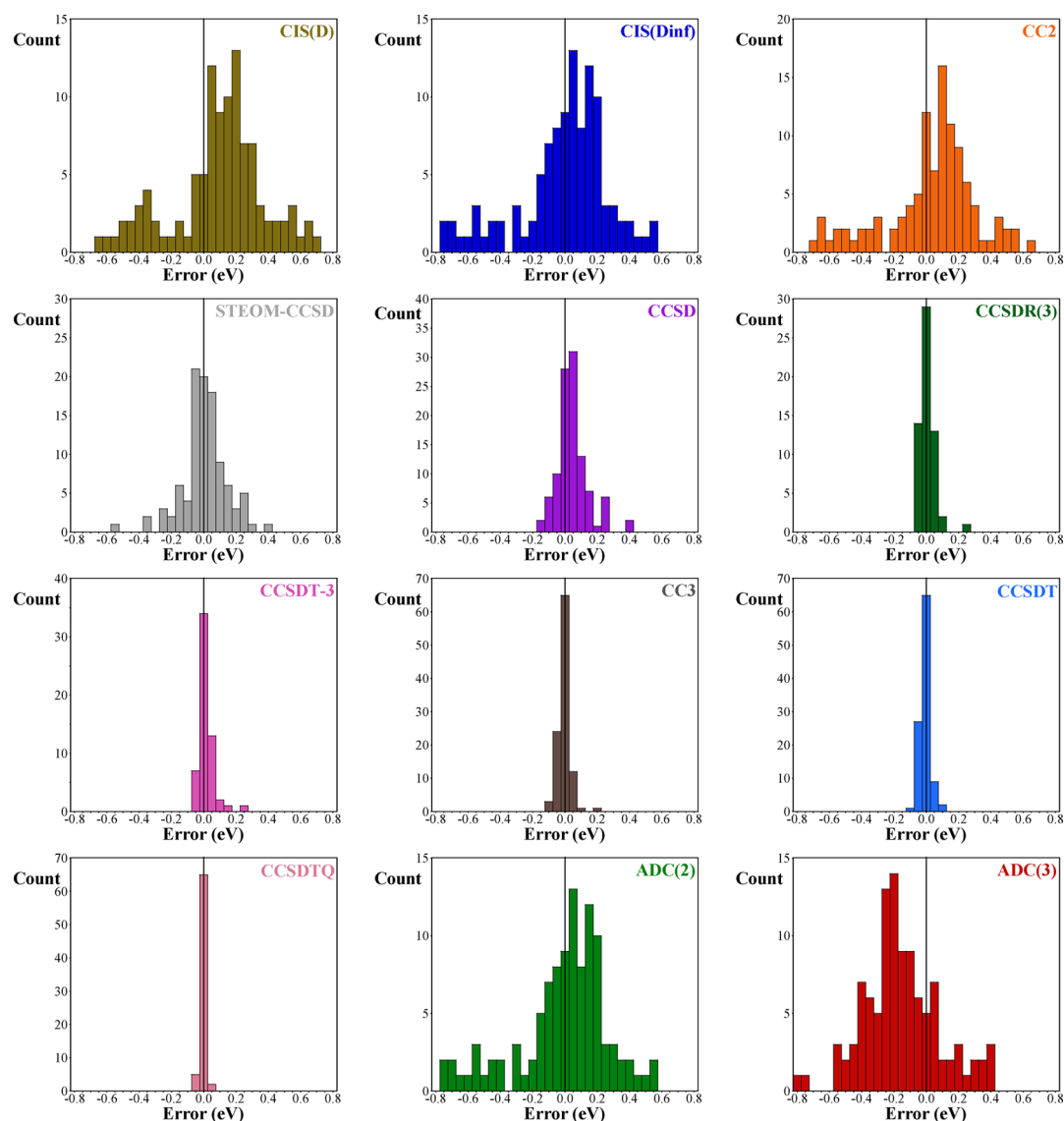


Figure 1. Histograms of the error patterns for several wave function methods compared to TBE(FC). Note the variation of scaling of the vertical axes.

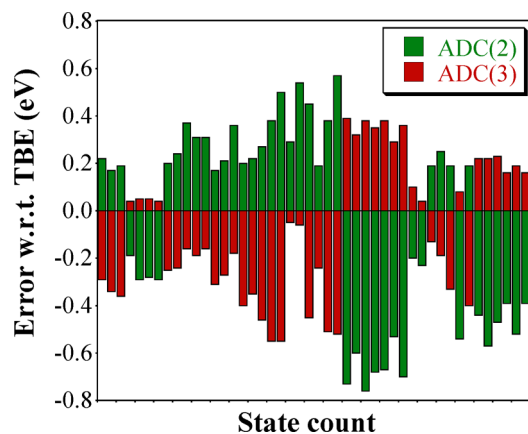


Figure 2. Comparison between the errors obtained with ADC(2) and ADC(3) [compared to TBE(FC)] for the 46 states for which ADC(2) yields an absolute deviation larger than 0.15 eV. All values are in eV.

CCSDTQ make the size of the atomic basis the central bottleneck. For this reason, we have tested the use of one of the most compact basis encompassing both diffuse and polarization functions, namely Pople's 6-31+G(d). We have performed CC3, CCSDT, and CCSDTQ calculations with this particular basis. The results are collected in the [Supporting Information](#) (Table S8). First, we compare the 6-31+G(d) results to those obtained with the same theoretical method in conjunction with the *aug-cc-pVTZ* basis set. As expected, large discrepancies are found with mean absolute deviation of 0.20, 0.19, and 0.25 eV, for CC3, CCSDT, and CCSDTQ, respectively.¹⁷⁵ Second, by adding the differences between the CC3/*aug-cc-pVTZ* and CC3/6-31+G(d) results to the CCSDT/6-31+G(d) and CCSDTQ/6-31+G(d) values, we obtained improved values. Such procedure yields very good estimates of the actual *aug-cc-pVTZ* results, as the MAE are down to 0.01 eV with no error larger than 0.04 eV for both CCSDT and CCSDTQ. This is a particularly remarkable result for Rydberg states that are extremely basis set dependent. For example, for the $^3A_2(n \rightarrow 3p)$ transition in water, the CCSDTQ/6-31+G(d) value of 10.34 eV is more than 1 eV

above its CCSDTQ/*aug-cc-pVTZ* counterpart (9.23 eV, see Table 1). Applying the CC3 basis set correction makes the final error as small as 0.03 eV. This composite methodology opens the way to calculations on larger systems without significant loss of accuracy.

4. CONCLUSIONS AND OUTLOOK

We have defined a set of more than 100 vertical transition energies, as close as possible to the FCI limit. To this end, we have used both the coupled cluster route up to the highest computationally possible order and the selected configuration interaction route up to the largest technically affordable number of determinants, that is here about few millions. These calculations have been performed on 18 compounds encompassing one, two or three non-hydrogen atoms, using geometries optimized at the CC3 level and a series of diffuse Dunning's basis sets of increasing size. It was certainly gratifying to find extremely good agreements between the results obtained independently with these two distinct approaches with typical differences as small as 0.01 eV between CCSDTQ and exFCI transition energies. In fact, during the course of this joint work, the two groups involved in this study were able to detect misprints or incorrect assignments in each others calculations even when the differences were apparently negligible. For the two diatomic molecules considered in this work, N₂ and CO, the mean absolute deviation between our theoretical best estimates and the "experimental" vertical transition energies deduced from spectroscopic measurements using a numerical solution of the nuclear Schrödinger equation is as small as 0.04 eV, and it was possible to resolve previous inconsistencies between these "experimental" values. A significant share of the remaining error is likely related to the use of theoretically determined geometries. Although, it is not possible to provide a definitive error bar for the 110 TBE listed in this work, our estimate, based on the differences between the two routes as well as the extrapolations used in the sCI procedure, is ± 0.03 eV.

In another part of this work, we have used the TBE(FC)/*aug-cc-pVTZ* values to benchmark a series of 12 popular wave function approaches. For the computationally most effective approaches, CIS(D), CIS(D_∞), ADC(2), and CC2, we found average deviations of ca. 0.21–0.25 eV with strong similarities between the ADC(2) and CC2 results. Both conclusions are backed up by previous works. Likewise, we obtained the expected trend that CCSD overestimates the transition energies, though with an amplitude that is quite small here, likely due to the small size of the compounds investigated. More interestingly, we could demonstrate that STEOM-CCSD is, on average, as accurate as CCSD, and we were also able to benchmark the methods including contributions from triples using reliable theoretical references. Interestingly, we found no significant differences between CCSDT-3, CC3, and CCSDT, which all yield a MAE of 0.03 eV. In other words, we could not demonstrate that CCSDT is statistically more accurate than its approximated (and computationally more effective) forms, nor highlight significant differences between CCSDT-3 and CC3. We have observed that the use of perturbative triples, as in CCSDR(3), allows to correct most of the CCSD error. This evidences that CCSDR(3) is a computationally appealing method as it gives average deviations only slightly larger than with iterative triples. In contrast, for the present set of molecules, ADC(3) was found significantly less accurate than CC3, and it was showed that ADC(3) overcorrects ADC(2).

Whether this surprising result is related to the size of the compounds or is a more general trend remains to be confirmed.

As stated several times throughout this work, the size of the considered molecules is certainly one of the main limitations of the present effort, as it introduces a significant bias, e.g., charge-transfer over several Å are totally absent of the set. Obviously, the respective $O(N^{10})$ and $O(e^N)$ formal scalings of CCSDTQ and FCI do not offer an easy pathway to circumvent this limit. Nevertheless, it appears that performing exFCI calculations with a relatively compact basis, e.g., *aug-cc-pVDZ* or even 6-31+G(d), and correcting the basis set effects with a more affordable approach, e.g., CC3, might be a valuable and efficient approach to reach accurate vertical excitations energies for larger molecules, at least for the electronic transitions presenting a dominant single excitation character. Indeed, we have shown here that such basis set extrapolation approach is trustworthy. We are currently hiking along that path.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.jctc.8b00406.

Basis set and frozen-core effects, geometries used, full list of transition energies for the benchmark section, additional statistical analysis, 6-31+G(d) results, and additional information for selected CI calculations (PDF)

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Notes

The authors declare no competing financial interest.

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■ REFERENCES

- (1) Hegarty, D.; Robb, M. A. Application of Unitary Group Methods to Configuration Interaction Calculations. *Mol. Phys.* **1979**, *38*, 1795–1812.
- (2) Taylor, P. R. Analytical MCSCF Energy Gradients: Treatment of Symmetry and CASSCF Applications to Propadienone. *J. Comput. Chem.* **1984**, *5*, 589–597.
- (3) Casida, M. E.; Huix-Rotllant, M. Progress in Time-Dependent Density-Functional Theory. *Annu. Rev. Phys. Chem.* **2012**, *63*, 287–323.
- (4) Ullrich, C. *Time-Dependent Density-Functional Theory: Concepts and Applications*; Oxford Graduate Texts; Oxford University Press: New York, 2012.

- (5) Andersson, K.; Malmqvist, P. A.; Roos, B. O.; Sadlej, A. J.; Wolinski, K. Second-Order Perturbation Theory With a CASSCF Reference Function. *J. Phys. Chem.* **1990**, *94*, 5483–5488.
- (6) Dreuw, A.; Wormit, M. The Algebraic Diagrammatic Construction Scheme for the Polarization Propagator for the Calculation of Excited States. *WIREs Comput. Mol. Sci.* **2015**, *5*, 82–95.
- (7) Christiansen, O.; Koch, H.; Jørgensen, P. The Second-Order Approximate Coupled Cluster Singles and Doubles Model CC2. *Chem. Phys. Lett.* **1995**, *243*, 409–418.
- (8) Hättig, C.; Weigend, F. CC2 Excitation Energy Calculations on Large Molecules Using the Resolution of the Identity Approximation. *J. Chem. Phys.* **2000**, *113*, 5154–5161.
- (9) Laurent, A. D.; Adamo, C.; Jacquemin, D. Dye Chemistry with Time-Dependent Density Functional Theory. *Phys. Chem. Chem. Phys.* **2014**, *16*, 14334–14356.
- (10) Winter, N. O. C.; Graf, N. K.; Leutwyler, S.; Hättig, C. Benchmarks for 0–0 Transitions of Aromatic Organic Molecules: DFT/B3LYP, ADC(2), CC2, SOS-CC2 and SCS-CC2 Compared to High-Resolution Gas-Phase Data. *Phys. Chem. Chem. Phys.* **2013**, *15*, 6623–6630.
- (11) Jacquemin, D.; Duchemin, I.; Blase, X. 0–0 Energies Using Hybrid Schemes: Benchmarks of TD-DFT, CIS(D), ADC(2), CC2 and BSE/GW formalisms for 80 Real-Life Compounds. *J. Chem. Theory Comput.* **2015**, *11*, 5340–5359.
- (12) Oruganti, B.; Fang, C.; Durbbeej, B. Assessment of a Composite CC2/DFT Procedure for Calculating 0–0 Excitation Energies of Organic Molecules. *Mol. Phys.* **2016**, *114*, 3448–3463.
- (13) Oddershede, J.; Grüner, N. E.; Diercks, G. H. Comparison Between Equation of Motion and Polarization Propagator Calculations. *Chem. Phys.* **1985**, *97*, 303–310.
- (14) Schwabe, T.; Goerigk, L. Time-Dependent Double-Hybrid Density Functionals with Spin-Component and Spin-Opposite Scaling. *J. Chem. Theory Comput.* **2017**, *13*, 4307–4323.
- (15) Christiansen, O.; Koch, H.; Jørgensen, P. Response Functions in the CC3 Iterative Triple Excitation Model. *J. Chem. Phys.* **1995**, *103*, 7429–7441.
- (16) Koch, H.; Christiansen, O.; Jørgensen, P.; Sanchez de Merás, A. M.; Helgaker, T. The CC3Model: An Iterative Coupled Cluster Approach Including Connected Triples. *J. Chem. Phys.* **1997**, *106*, 1808–1818.
- (17) Leang, S. S.; Zahariev, F.; Gordon, M. S. Benchmarking the Performance of Time-Dependent Density Functional Methods. *J. Chem. Phys.* **2012**, *136*, 104101.
- (18) Parac, M.; Grimme, S. Comparison of Multireference Möller-Plesset Theory and Time-Dependent Methods for the Calculation of Vertical Excitation Energies of Molecules. *J. Phys. Chem. A* **2002**, *106*, 6844–6850.
- (19) Dierksen, M.; Grimme, S. The Vibronic Structure of Electronic Absorption Spectra of Large Molecules: A Time-Dependent Density Functional Study on the Influence of Exact Hartree-Fock Exchange. *J. Phys. Chem. A* **2004**, *108*, 10225–10237.
- (20) Grimme, S.; Izgorodina, E. I. Calculation of 0–0 Excitation Energies of Organic Molecules by CIS(D) Quantum Chemical Methods. *Chem. Phys.* **2004**, *305*, 223–230.
- (21) Rhee, Y. M.; Head-Gordon, M. Scaled Second-Order Perturbation Corrections to Configuration Interaction Singles: Efficient and Reliable Excitation Energy Methods. *J. Phys. Chem. A* **2007**, *111*, 5314–5326.
- (22) Peach, M. J. G.; Benfield, P.; Helgaker, T.; Tozer, D. J. Excitation Energies in Density Functional Theory: an Evaluation and a Diagnostic Test. *J. Chem. Phys.* **2008**, *128*, 044118.
- (23) Jacquemin, D.; Perpète, E. A.; Scuseria, G. E.; Ciofini, I.; Adamo, C. TD-DFT Performance for the Visible Absorption Spectra of Organic Dyes: Conventional Versus Long-Range Hybrids. *J. Chem. Theory Comput.* **2008**, *4*, 123–135.
- (24) Jacquemin, D.; Wathelot, V.; Perpète, E. A.; Adamo, C. Extensive TD-DFT Benchmark: Singlet-Excited States of Organic Molecules. *J. Chem. Theory Comput.* **2009**, *5*, 2420–2435.
- (25) Goerigk, L.; Moellmann, J.; Grimme, S. Computation of Accurate Excitation Energies for Large Organic Molecules with Double-Hybrid Density Functionals. *Phys. Chem. Chem. Phys.* **2009**, *11*, 4611–4620.
- (26) Caricato, M.; Trucks, G. W.; Frisch, M. J.; Wiberg, K. B. Electronic Transition Energies: A Study of the Performance of a Large Range of Single Reference Density Functional and Wave Function Methods on Valence and Rydberg States Compared to Experiment. *J. Chem. Theory Comput.* **2010**, *6*, 370–383.
- (27) Jacquemin, D.; Planchat, A.; Adamo, C.; Mennucci, B. A TD-DFT Assessment of Functionals for Optical 0–0 Transitions in Solvated Dyes. *J. Chem. Theory Comput.* **2012**, *8*, 2359–2372.
- (28) Isegawa, M.; Peverati, R.; Truhlar, D. G. Performance of Recent and High-Performance Approximate Density Functionals for Time-Dependent Density Functional Theory Calculations of Valence and Rydberg Electronic Transition Energies. *J. Chem. Phys.* **2012**, *137*, 244104.
- (29) Hoyer, C. E.; Ghosh, S.; Truhlar, D. G.; Gagliardi, L. Multiconfiguration Pair-Density Functional Theory Is as Accurate as CASPT2 for Electronic Excitation. *J. Phys. Chem. Lett.* **2016**, *7*, 586–591.
- (30) Schreiber, M.; Silva-Junior, M. R.; Sauer, S. P. A.; Thiel, W. Benchmarks for Electronically Excited States: CASPT2, CC2, CCSD and CC3. *J. Chem. Phys.* **2008**, *128*, 134110.
- (31) Silva-Junior, M. R.; Sauer, S. P. A.; Schreiber, M.; Thiel, W. Basis Set Effects on Coupled Cluster Benchmarks of Electronically Excited States: CC3, CCSDR(3) and CC2. *Mol. Phys.* **2010**, *108*, 453–465.
- (32) Silva-Junior, M. R.; Schreiber, M.; Sauer, S. P. A.; Thiel, W. Benchmarks of Electronically Excited States: Basis Set Effects. Benchmarks of Electronically Excited States: Basis Set Effects on CASPT2 Results. *J. Chem. Phys.* **2010**, *133*, 174318.
- (33) Silva-Junior, M. R.; Thiel, W. Benchmark of Electronically Excited States for Semiempirical Methods: MNDO, AM1, PM3, OM1, OM2, OM3, INDO/S, and INDO/S2. *J. Chem. Theory Comput.* **2010**, *6*, 1546–1564.
- (34) Domínguez, A.; Aradi, B.; Frauenheim, T.; Lutsker, V.; Niehaus, T. A. Extensions of the Time-Dependent Density Functional Based Tight-Binding Approach. *J. Chem. Theory Comput.* **2013**, *9*, 4901–4914.
- (35) Voityuk, A. A. INDO/X: A New Semiempirical Method for Excited States of Organic and Biological Molecules. *J. Chem. Theory Comput.* **2014**, *10*, 4950–4958.
- (36) Silva-Junior, M. R.; Schreiber, M.; Sauer, S. P. A.; Thiel, W. Benchmarks for Electronically Excited States: Time-Dependent Density Functional Theory and Density Functional Theory Based Multireference Configuration Interaction. *J. Chem. Phys.* **2008**, *129*, 104103.
- (37) Rohrdanz, M. A.; Martins, K. M.; Herbert, J. M. A Long-Range-Corrected Density Functional That Performs Well for Both Ground-State Properties and Time-Dependent Density Functional Theory Excitation Energies, Including Charge-Transfer Excited States. *J. Chem. Phys.* **2009**, *130*, 054112.
- (38) Jacquemin, D.; Perpète, E. A.; Ciofini, I.; Adamo, C. Assessment of Functionals for TD-DFT Calculations of Singlet-Triplet Transitions. *J. Chem. Theory Comput.* **2010**, *6*, 1532–1537.
- (39) Jacquemin, D.; Perpète, E. A.; Ciofini, I.; Adamo, C.; Valero, R.; Zhao, Y.; Truhlar, D. G. On the Performances of the M06 Family of Density Functionals for Electronic Excitation Energies. *J. Chem. Theory Comput.* **2010**, *6*, 2071–2085.
- (40) Mardirossian, N.; Parkhill, J. A.; Head-Gordon, M. Benchmark Results for Empirical Post-GGA Functionals: Difficult Exchange Problems and Independent Tests. *Phys. Chem. Chem. Phys.* **2011**, *13*, 19325–19337.
- (41) Jacquemin, D.; Perpète, E. A.; Ciofini, I.; Adamo, C. Assessment of the ω B97 Family for Excited-State Calculations. *Theor. Chem. Acc.* **2011**, *128*, 127–136.
- (42) Huix-Rotllant, M.; Ipatov, A.; Rubio, A.; Casida, M. E. Assessment of Dressed Time-Dependent Density-Functional Theory

for the Low-Lying Valence States of 28 Organic Chromophores. *Chem. Phys.* **2011**, *391*, 120–129.

(43) Della Sala, F.; Fabiano, E. Accurate Singlet and Triplet Excitation Energies Using the Localized Hartree-Fock Kohn-Sham Potential. *Chem. Phys.* **2011**, *391*, 19–26.

(44) Trani, F.; Scalmani, G.; Zheng, G. S.; Camimeo, I.; Frisch, M. J.; Barone, V. Time-Dependent Density Functional Tight Binding: New Formulation and Benchmark of Excited States. *J. Chem. Theory Comput.* **2011**, *7*, 3304–3313.

(45) Peverati, R.; Truhlar, D. G. Performance of the M11 and M11-L Density Functionals for Calculations of Electronic Excitation Energies by Adiabatic Time-Dependent Density Functional Theory. *Phys. Chem. Chem. Phys.* **2012**, *14*, 11363–11370.

(46) Maier, T. M.; Bahmann, H.; Arbuznikov, A. V.; Kaupp, M. Validation of Local Hybrid Functionals for TDDFT Calculations of Electronic Excitation Energies. *J. Chem. Phys.* **2016**, *144*, 074106.

(47) Sauer, S. P.; Pitzner-Frydendahl, H. F.; Buse, M.; Jensen, H. J. A.; Thiel, W. Performance of SOPPA-Based Methods in the Calculation of Vertical Excitation Energies and Oscillator Strengths. *Mol. Phys.* **2015**, *113*, 2026–2045.

(48) Harbach, P. H. P.; Wormit, M.; Dreuw, A. The Third-Order Algebraic Diagrammatic Construction Method (ADC(3)) for the Polarization Propagator for Closed-Shell Molecules: Efficient Implementation and Benchmarking. *J. Chem. Phys.* **2014**, *141*, 064113.

(49) Schapiro, I.; Sivalingam, K.; Neese, F. Assessment of *n*-Electron Valence State Perturbation Theory for Vertical Excitation Energies. *J. Chem. Theory Comput.* **2013**, *9*, 3567–3580.

(50) Yang, Y.; Peng, D.; Lu, J.; Yang, W. Excitation Energies from Particle-Particle Random Phase Approximation: Davidson Algorithm and Benchmark Studies. *J. Chem. Phys.* **2014**, *141*, 124104.

(51) Sauer, S. P. A.; Schreiber, M.; Silva-Junior, M. R.; Thiel, W. Benchmarks for Electronically Excited States: A Comparison of Noniterative and Iterative Triples Corrections in Linear Response Coupled Cluster Methods: CCSDR(3) versus CC3. *J. Chem. Theory Comput.* **2009**, *5*, 555–564.

(52) Demel, O.; Datta, D.; Nooijen, M. Additional Global Internal Contraction in Variations of Multireference Equation of Motion Coupled Cluster Theory. *J. Chem. Phys.* **2013**, *138*, 134108.

(53) Piecuch, P.; Hansen, J. A.; Ajala, A. O. Benchmarking the Completely Renormalised Equation-Of-Motion Coupled-Cluster Approaches for Vertical Excitation Energies. *Mol. Phys.* **2015**, *113*, 3085–3127.

(54) Tajti, A.; Szalay, P. G. Investigation of the Impact of Different Terms in the Second Order Hamiltonian on Excitation Energies of Valence and Rydberg States. *J. Chem. Theory Comput.* **2016**, *12*, 5477–5482.

(55) Rishi, V.; Perera, A.; Nooijen, M.; Bartlett, R. J. Excited States from Modified Coupled Cluster Methods: Are They Any Better Than EOM CCSD? *J. Chem. Phys.* **2017**, *146*, 144104.

(56) Dutta, A. K.; Nooijen, M.; Neese, F.; Izsák, R. Exploring the Accuracy of a Low Scaling Similarity Transformed Equation of Motion Method for Vertical Excitation Energies. *J. Chem. Theory Comput.* **2018**, *14*, 72–91.

(57) Watson, T. J.; Lotrich, V. F.; Szalay, P. G.; Perera, A.; Bartlett, R. J. Benchmarking for Perturbative Triple-Excitations in EE-EOM-CC Methods. *J. Phys. Chem. A* **2013**, *117*, 2569–2579.

(58) Kánnár, D.; Szalay, P. G. Benchmarking Coupled Cluster Methods on Valence Singlet Excited States. *J. Chem. Theory Comput.* **2014**, *10*, 3757–3765.

(59) Kánnár, D.; Tajti, A.; Szalay, P. G. Accuracy of Coupled Cluster Excitation Energies in Diffuse Basis Sets. *J. Chem. Theory Comput.* **2017**, *13*, 202–209.

(60) Laurent, A.; Blondel, A.; Jacquemin, D. Choosing an Atomic Basis Set for TD-DFT, SOPPA, ADC(2), CIS(D), CC2 and EOM-CCSD Calculations of Low-Lying Excited States of Organic Dyes. *Theor. Chem. Acc.* **2015**, *134*, 76.

(61) Budzák, Š.; Scalmani, G.; Jacquemin, D. Accurate Excited-State Geometries: a CASPT2 and Coupled-Cluster Reference Database for Small Molecules. *J. Chem. Theory Comput.* **2017**, *13*, 6237–6252.

(62) Purvis, G. P., III; Bartlett, R. J. A Full Coupled-Cluster Singles and Doubles Model: The Inclusion of Disconnected Triples. *J. Chem. Phys.* **1982**, *76*, 1910–1918.

(63) Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Scalmani, G.; Barone, V.; Petersson, G. A.; Nakatsuji, H.; Li, X.; Caricato, M.; Marenich, A. V.; Bloino, J.; Janesko, B. G.; Gomperts, R.; Mennucci, B.; Hratchian, H. P.; Ortiz, J. V.; Izmaylov, A. F.; Sonnenberg, J. L.; Williams-Young, D.; Ding, F.; Lipparini, F.; Egidi, F.; Goings, J.; Peng, B.; Petrone, A.; Henderson, T.; Ranasinghe, D.; Zakrzewski, V. G.; Gao, J.; Rega, N.; Zheng, G.; Liang, W.; Hada, M.; Ehara, M.; Toyota, K.; Fukuda, R.; Hasegawa, J.; Ishida, M.; Nakajima, T.; Honda, Y.; Kitao, O.; Nakai, H.; Vreven, T.; Throssell, K.; Montgomery, J. A., Jr.; Peralta, J. E.; Ogliaro, F.; Bearpark, M. J.; Heyd, J. J.; Brothers, E. N.; Kudin, K. N.; Staroverov, V. N.; Keith, T. A.; Kobayashi, R.; Normand, J.; Raghavachari, K.; Rendell, A. P.; Burant, J. C.; Iyengar, S. S.; Tomasi, J.; Cossi, M.; Millam, J. M.; Klene, M.; Adamo, C.; Cammi, R.; Ochterski, J. W.; Martin, R. L.; Morokuma, K.; Farkas, O.; Foresman, J. B.; Fox, D. J. *Gaussian 16*, Revision A.03; Gaussian Inc.: Wallingford, CT, 2016.

(64) For formamide, the CCSD/def2-TZVPP vibrational frequency calculation returns one imaginary frequency, which disappears at the CCSD/aug-cc-pVTZ level.

(65) Aidas, K.; Angeli, C.; Bak, K. L.; Bakken, V.; Bast, R.; Boman, L.; Christiansen, O.; Cimiraglia, R.; Coriani, S.; Dahle, P.; Dalskov, E. K.; Ekström, U.; Enevoldsen, T.; Eriksen, J. J.; Ettenhuber, P.; Fernández, B.; Ferrighi, L.; Flieg, H.; Frediani, L.; Hald, K.; Halkier, A.; Hättig, C.; Heiberg, H.; Helgaker, T.; Hennum, A. C.; Hettner, H.; Hjertenæs, E.; Høst, S.; Høyvik, I.-M.; Iozzi, M. F.; Jansík, B.; Jensen, H. J. A.; Jonsson, D.; Jørgensen, P.; Kauczor, J.; Kirpekar, S.; Kjærgaard, T.; Klopper, W.; Knecht, S.; Kobayashi, R.; Koch, H.; Kongsted, J.; Krapp, A.; Kristensen, K.; Ligabue, A.; Lutnæs, O. B.; Melo, J. I.; Mikkelsen, K. V.; Myhre, R. H.; Neiss, C.; Nielsen, C. B.; Norman, P.; Olsen, J.; Olsen, J. M. H.; Osted, A.; Packer, M. J.; Pawłowski, F.; Pedersen, T. B.; Provasi, P. F.; Reine, S.; Rinkevicius, Z.; Ruden, T. A.; Ruud, K.; Rybkin, V. V.; Salek, P.; Samson, C. C. M.; de Merás, A. S.; Saue, T.; Sauer, S. P. A.; Schimmelpfennig, B.; Sneskov, K.; Steindal, A. H.; Sylvester-Hvid, K. O.; Taylor, P. R.; Teale, A. M.; Tellgren, E. I.; Tew, D. P.; Thorvaldsen, A. J.; Thøgersen, L.; Vahtras, O.; Watson, M. A.; Wilson, D. J. D.; Ziolkowski, M.; Ågren, H. The Dalton Quantum Chemistry Program System. *WIREs Comput. Mol. Sci.* **2014**, *4*, 269–284.

(66) Stanton, J. F.; Gauss, J.; Cheng, L.; Harding, M. E.; Matthews, D. A.; Szalay, P. G. *CFOUR, Coupled-Cluster techniques for Computational Chemistry, a quantum-chemical program package*; with contributions from A. A. Auer, R. J. Bartlett, U. Benedikt, C. Berger, D. E. Bernholdt, Y. J. Bomble, O. Christiansen, F. Engel, R. Faber, M. Heckert, O. Heun, M. Hilgenberg, C. Huber, T.-C. Jagau, D. Jonsson, J. Jusélius, T. Kirsch, K. Klein, W. J. Lauderdale, F. Lipparini, T. Metzroth, L. A. Mück, D. P. O'Neill, D. R. Price, E. Prochnow, C. Puzzarini, K. Ruud, F. Schiffrmann, W. Schwabach, C. Simmons, S. Stopkowicz, A. Tajti, J. Vázquez, F. Wang, J. D. Watts and the integral packages MOLECULE (J. Almlöf and P. R. Taylor), PROPS (P. R. Taylor), ABACUS (T. Helgaker, H. J. Aa. Jensen, P. Jørgensen, and J. Olsen), and ECP routines by A. V. Mitin and C. van Wüllen. For the current version, see <http://www.cfour.de>.

(67) Kállay, M.; Gauss, J. Calculation of Excited-State Properties Using General Coupled-Cluster and Configuration-Interaction Models. *J. Chem. Phys.* **2004**, *121*, 9257–9269.

(68) Neese, F. The ORCA Program System. *WIREs Comput. Mol. Sci.* **2012**, *2*, 73–78.

(69) Rolik, Z.; Szegedy, L.; Ladjánszki, I.; Ladóczki, B.; Kállay, M. An Efficient Linear-Scaling CCSD(T) Method Based on Local Natural Orbitals. *J. Chem. Phys.* **2013**, *139*, 094105.

(70) Kállay, M.; Rolik, Z.; Csontos, J.; Nagy, P.; Samu, G.; Mester, D.; Csóka, J.; Szabó, B.; Ladjánszki, I.; Szegedy, L.; Ladóczki, B.;

Petrov, K.; Farkas, M.; Mezei, P. D.; Hégyely, B. MRCC, Quantum Chemical Program. 2017; See: www.mrcc.hu.

(71) Shao, Y.; Gan, Z.; Epifanovsky, E.; Gilbert, A. T.; Wormit, M.; Kussmann, J.; Lange, A. W.; Behn, A.; Deng, J.; Feng, X.; Ghosh, D.; Goldey, M.; Horn, P. R.; Jacobson, L. D.; Kaliman, I.; Khaliullin, R. Z.; Kus, T.; Landau, A.; Liu, J.; Proynov, E. I.; Rhee, Y. M.; Richard, R. M.; Rohrdanz, M. A.; Steele, R. P.; Sundstrom, E. J.; Woodcock, H. L.; Zimmerman, P. M.; Zuev, D.; Albrecht, B.; Alguire, E.; Austin, B.; Beran, G. J. O.; Bernard, Y. A.; Berquist, E.; Brandhorst, K.; Bravaya, K. B.; Brown, S. T.; Casanova, D.; Chang, C.-M.; Chen, Y.; Chien, S. H.; Closser, K. D.; Crittenden, D. L.; Diedenhofen, M.; DiStasio, R. A.; Do, H.; Dutoi, A. D.; Edgar, R. G.; Fatehi, S.; Fusti-Molnar, L.; Ghysels, A.; Golubeva-Zadorozhnaya, A.; Gomes, J.; Hanson-Heine, M. W.; Harbach, P. H.; Hauser, A. W.; Hohenstein, E. G.; Holden, Z. C.; Jagau, T.-C.; Ji, H.; Kaduk, B.; Khistyayev, K.; Kim, J.; Kim, J.; King, R. A.; Klunzinger, P.; Kosenkov, D.; Kowalczyk, T.; Krauter, C. M.; Lao, K. U.; Laurent, A. D.; Lawler, K. V.; Levchenko, S. V.; Lin, C. Y.; Liu, F.; Livshits, E.; Lochan, R. C.; Luenser, A.; Manohar, P.; Manzer, S. F.; Mao, S.-P.; Mardirossian, N.; Marenich, A. V.; Maurer, S. A.; Mayhall, N. J.; Neuscamman, E.; Oana, C. M.; Olivares-Amaya, R.; O'Neill, D. P.; Parkhill, J. A.; Perrine, T. M.; Peverati, R.; Prociuk, A.; Rehn, D. R.; Rosta, E.; Russ, N. J.; Sharada, S. M.; Sharma, S.; Small, D. W.; Sodt, A.; Stein, T.; Stück, D.; Su, Y.-C.; Thom, A. J.; Tsuchimochi, T.; Vanovschi, V.; Vogt, L.; Vydrov, O.; Wang, T.; Watson, M. A.; Wenzel, J.; White, A.; Williams, C. F.; Yang, J.; Yeganeh, S.; Yost, S. R.; You, Z.-Q.; Zhang, I. Y.; Zhang, X.; Zhao, Y.; Ziegler, B. R.; Chan, G. K.; Chipman, D. M.; Cramer, C. J.; Goddard, W. A.; Gordon, M. S.; Hehre, W. J.; Klamt, A.; Schaefer, H. F.; Schmidt, M. W.; Sherrill, C. D.; Truhlar, D. G.; Warshel, A.; Xu, X.; Aspuru-Guzik, A.; Baer, R.; Bell, A. T.; Besley, N. A.; Chai, J.-D.; Dreuw, A.; Dunietz, B. D.; Furlani, T. R.; Gwaltney, S. R.; Hsu, C.-P.; Jung, Y.; Kong, J.; Lambrecht, D. S.; Liang, W.; Ochsenfeld, C.; Rassolov, V. A.; Slipchenko, L. V.; Subotnik, J. E.; Van Voorhis, T.; Herbert, J. M.; Krylov, A. I.; Gill, P. M.; Head-Gordon, M. Advances in Molecular Quantum Chemistry Contained in the Q-Chem 4 Program Package. *Mol. Phys.* **2015**, *113*, 184–215.

(72) Watts, J. D.; Bartlett, R. J. Iterative and Non-Iterative Triple Excitation Corrections in Coupled-Cluster Methods for Excited Electronic States: the EOM-CCSDT-3 and EOM-CCSD(\bar{T}) Methods. *Chem. Phys. Lett.* **1996**, *258*, 581–588.

(73) Prochnow, E.; Harding, M. E.; Gauss, J. Parallel Calculation of CCSDT and Mk-MRCCSDT Energies. *J. Chem. Theory Comput.* **2010**, *6*, 2339–2347.

(74) Noga, J.; Bartlett, R. J. The Full CCSDT Model for Molecular Electronic Structure. *J. Chem. Phys.* **1987**, *86*, 7041–7050.

(75) Head-Gordon, M.; Rico, R. J.; Oumi, M.; Lee, T. J. A Doubles Correction to Electronic Excited States From Configuration Interaction in the Space of Single Substitutions. *Chem. Phys. Lett.* **1994**, *219*, 21–29.

(76) Head-Gordon, M.; Maurice, D.; Oumi, M. A Perturbative Correction to Restricted Open-Shell Configuration-Interaction with Single Substitutions for Excited-States of Radicals. *Chem. Phys. Lett.* **1995**, *246*, 114–121.

(77) Christiansen, O.; Koch, H.; Jørgensen, P. Perturbative Triple Excitation Corrections to Coupled Cluster Singles and Doubles Excitation Energies. *J. Chem. Phys.* **1996**, *105*, 1451–1459.

(78) Nooijen, M.; Bartlett, R. J. A New Method for Excited States: Similarity Transformed Equation-Of-Motion Coupled-Cluster Theory. *J. Chem. Phys.* **1997**, *106*, 6441–6448.

(79) Head-Gordon, M.; Oumi, M.; Maurice, D. Quasidegenerate Second-Order Perturbation Corrections to Single-Excitation Configuration Interaction. *Mol. Phys.* **1999**, *96*, 593–602.

(80) Kucharski, S. A.; Bartlett, R. J. Recursive Intermediate Factorization and Complete Computational Linearization of the Coupled-Cluster Single, Double, Triple, and Quadruple Excitation Equations. *Theor. Chim. Acta* **1991**, *80*, 387–405.

(81) Bender, C. F.; Davidson, E. R. Studies in Configuration Interaction: The First-Row Diatomic Hydrides. *Phys. Rev.* **1969**, *183*, 23–30.

(82) Whitten, J. L.; Hackmeyer, M. Configuration Interaction Studies of Ground and Excited States of Polyatomic Molecules. I. The CI Formulation and Studies of Formaldehyde. *J. Chem. Phys.* **1969**, *51*, 5584–5596.

(83) Giner, E.; Scemama, A.; Caffarel, M. Using Perturbatively Selected Configuration Interaction in Quantum Monte Carlo Calculations. *Can. J. Chem.* **2013**, *91*, 879–885.

(84) Caffarel, M.; Giner, E.; Scemama, A.; Ramírez-Solís, A. Spin Density Distribution in Open-Shell Transition Metal Systems: A Comparative Post-Hartree-Fock, Density Functional Theory, and Quantum Monte Carlo Study of the CuCl₂ Molecule. *J. Chem. Theory Comput.* **2014**, *10*, 5286–5296.

(85) Giner, E.; Scemama, A.; Caffarel, M. Fixed-Node Diffusion Monte Carlo Potential Energy Curve of the Fluorine Molecule F₂ Using Selected Configuration Interaction Trial Wavefunctions. *J. Chem. Phys.* **2015**, *142*, 044115.

(86) Garniron, Y.; Giner, E.; Malrieu, J.-P.; Scemama, A. Alternative Definition of Excitation Amplitudes in Multi-Reference State-Specific Coupled Cluster. *J. Chem. Phys.* **2017**, *146*, 154107.

(87) Caffarel, M.; Applencourt, T.; Giner, E.; Scemama, A. Toward an Improved Control of the Fixed-Node Error in Quantum Monte Carlo: The Case of the Water Molecule. *J. Chem. Phys.* **2016**, *144*, 151103.

(88) Holmes, A. A.; Tubman, N. M.; Umrigar, C. J. Heat-Bath Configuration Interaction: An Efficient Selected Configuration Interaction Algorithm Inspired by Heat-Bath Sampling. *J. Chem. Theory Comput.* **2016**, *12*, 3674–3680.

(89) Sharma, S.; Holmes, A. A.; Jeanmairet, G.; Alavi, A.; Umrigar, C. J. Semistochastic Heat-Bath Configuration Interaction Method: Selected Configuration Interaction with Semistochastic Perturbation Theory. *J. Chem. Theory Comput.* **2017**, *13*, 1595–1604.

(90) Holmes, A. A.; Umrigar, C. J.; Sharma, S. Excited States Using Semistochastic Heat-Bath Configuration Interaction. *J. Chem. Phys.* **2017**, *147*, 164111.

(91) Chien, A. D.; Holmes, A. A.; Otten, M.; Umrigar, C. J.; Sharma, S.; Zimmerman, P. M. Excited States of Methylene, Polyenes, and Ozone from Heat-Bath Configuration Interaction. *J. Phys. Chem. A* **2018**, *122*, 2714–2722.

(92) Scemama, A.; Garniron, Y.; Caffarel, M.; Loos, P. F. Deterministic Construction of Nodal Surfaces Within Quantum Monte Carlo: The Case of FeS. *J. Chem. Theory Comput.* **2018**, *14*, 1395–1402.

(93) Huron, B.; Malrieu, J. P.; Rancurel, P. Iterative Perturbation Calculations of Ground and Excited State Energies from Multi-configurational Zeroth-Order Wavefunctions. *J. Chem. Phys.* **1973**, *58*, 5745–5759.

(94) Evangelisti, S.; Daudey, J.-P.; Malrieu, J.-P. Convergence of an Improved CIPSI Algorithm. *Chem. Phys.* **1983**, *75*, 91–102.

(95) Caffarel, M.; Applencourt, T.; Giner, E.; Scemama, A. ACS Symp. Ser. **2016**, *1234*, 15–46.

(96) Garniron, Y.; Scemama, A.; Loos, P.-F.; Caffarel, M. Hybrid Stochastic-Deterministic Calculation of the Second-Order Perturbative Contribution of Multireference Perturbation Theory. *J. Chem. Phys.* **2017**, *147*, 034101.

(97) Scemama, A.; Applencourt, T.; Garniron, Y.; Giner, E.; David, G.; Caffarel, M. *Quantum Package*, v1.0; 2016; https://github.com/LCPQ/quantum_package, https://github.com/LCPQ/quantum_package.

(98) Ralphs, K.; Serna, G.; Hargreaves, L. R.; Khakoo, M. A.; Winstead, C.; McKoy, V. Excitation of the Six Lowest Electronic Transitions in Water by 9–20 eV Electrons. *J. Phys. B: At., Mol. Opt. Phys.* **2013**, *46*, 125201.

(99) Cai, Z.-L.; Tozer, D. J.; Reimers, J. R. Time-Dependent Density-Functional Determination of Arbitrary Singlet and Triplet Excited-State Potential Energy Surfaces: Application to the Water Molecule. *J. Chem. Phys.* **2000**, *113*, 7084–7096.

(100) Li, X.; Paldus, J. General-Model-Space State-Universal Coupled-Cluster Method: Excitation Energies of Water. *Mol. Phys.* **2006**, *104*, 661–676.

- (101) Masuko, H.; Morioka, Y.; Nakamura, M.; Ishiguro, E.; Sasanuma, M. Absorption Spectrum of the H₂S Molecule In The Vacuum Ultraviolet Region. *Can. J. Phys.* **1979**, *57*, 745–760.
- (102) Abuain, T.; Walker, I. C.; Dance, D. F. Electronic Excitation of Hydrogen Sulphide by Impact With (Near-Threshold) Low-Energy Electrons. *J. Chem. Soc., Faraday Trans. 2* **1986**, *82*, 811–816.
- (103) Páleníková, J.; Kraus, M.; Neogrády, P.; Kellö, V.; Urban, M. Theoretical Study of Molecular Properties of Low-Lying Electronic Excited States of H₂O and H₂S. *Mol. Phys.* **2008**, *106*, 2333–2344.
- (104) Skerbele, A.; Lassetre, E. N. Electron-Impact Spectra. *J. Chem. Phys.* **1965**, *42*, 395–401.
- (105) Harshbarger, W. R. Identification of the \tilde{C} State of Ammonia by Electron Impact Spectroscopy. *J. Chem. Phys.* **1971**, *54*, 2504–2509.
- (106) Bartlett, R. J.; Del Bene, J. E.; Perera, S.; Mattie, R. Ammonia: The Prototypical Lone Pair Molecule. *J. Mol. Struct.: THEOCHEM* **1997**, *400*, 157–168.
- (107) Arfa, M. B.; Tronc, M. Lowest Energy Triplet States of Group Vb Hydrides: NH₃ (ND₃) and PH₃. *Chem. Phys.* **1991**, *155*, 143–148.
- (108) Abuain, T.; Walker, I. C.; Dance, D. F. The Lowest Triplet State in Ammonia and Methylamine Detected by Electron-Impact Excitation. *J. Chem. Soc., Faraday Trans. 2* **1984**, *80*, 641–645.
- (109) Rubio, M.; Serrano-Andrés, L.; Merchán, M. Excited States of the Water Molecule: Analysis of the Valence and Rydberg Character. *J. Chem. Phys.* **2008**, *128*, 104305.
- (110) Jacquemin, D.; Duchemin, I.; Blase, X. Benchmarking the Bethe-Salpeter Formalism on a Standard Organic Molecular Set. *J. Chem. Theory Comput.* **2015**, *11*, 3290–3304.
- (111) Chutjian, A.; Hall, R. I.; Trajmar, S. Electron-Impact Excitation of H₂O and D₂O at Various Scattering Angles and Impact Energies in the Energy-Loss Range 4.2–12 eV. *J. Chem. Phys.* **1975**, *63*, 892–898.
- (112) Pitarch-Ruiz, J.; Sánchez-Marín, J.; Martín, I.; Velasco, A. M. Vertical Excitation Energies and Ionization Potentials of H₂S. A Size-Consistent Self-Consistent Singles and Doubles Configuration Interaction (SC)²-MR-SDCI Calculation. *J. Phys. Chem. A* **2002**, *106*, 6508–6514.
- (113) Velasco, A. M.; Martín, I.; Pitarch-Ruiz, J.; Sánchez-Marín, J. MRSDCI Vertical Excitation Energies and MQDO Intensities for Electronic Transitions to Rydberg States in H₂S. *J. Phys. Chem. A* **2004**, *108*, 6724–6729.
- (114) Gupta, M.; Baluja, K. L. Application of R-Matrix Method to Electron-H₂S Collisions in the Low Energy Range. *Eur. Phys. J. D* **2007**, *41*, 475–483.
- (115) O'Brien Lantz, K.; Vaida, V. Direct Absorption Spectroscopy of the First Excited Electronic Band of Jet-Cooled H₂S. *Chem. Phys. Lett.* **1993**, *215*, 329–335.
- (116) Robin, M. B. In *Higher Excited States of Polyatomic Molecules*; Robin, M. B., Ed.; Academic Press: 1985; Vol. III.
- (117) Chantranupong, L.; Hirsch, G.; Buenker, R. J.; Kimura, M.; Dillon, M. A. Theoretical Study of the Electronic Spectrum of Ammonia: Generalized Oscillator Strength Calculations for the A-X Transition. *Chem. Phys.* **1991**, *154*, 13–21.
- (118) Huber, K. P.; Herzberg, G. *Constants of Diatomic Molecules; Molecular Spectra and Molecular Structure*; Van Nostrand: Princeton, NJ, 1979; Vol. 4.
- (119) Ben-Shlomo, S. B.; Kaldor, U. N₂ Excitations Below 15 eV by the Multireference Coupled-Cluster Method. *J. Chem. Phys.* **1990**, *92*, 3680–3682.
- (120) Nielsen, E. S.; Jorgensen, P.; Oddershede, J. Transition Moments and Dynamic Polarizabilities in a Second Order Polarization Propagator Approach. *J. Chem. Phys.* **1980**, *73*, 6238–6246.
- (121) Kucharski, S. A.; Wloch, M.; Musiał, M.; Bartlett, R. J. Coupled-Cluster Theory for Excited Electronic States: The Full Equation-Of-Motion Coupled-Cluster Single, Double, and Triple Excitation Method. *J. Chem. Phys.* **2001**, *115*, 8263–8266.
- (122) Dora, A.; Tennyson, J.; Chakrabarti, K. Higher Lying Resonances in Low-Energy Electron Scattering with Carbon Monoxide. *Eur. Phys. J. D* **2016**, *70*, 197.
- (123) Stahel, D.; Leoni, M.; Dressler, K. Nonadiabatic Representations of the $^1\Sigma_u^+$ and $^1\Pi_u$ States of the N₂ Molecule. *J. Chem. Phys.* **1983**, *79*, 2541–2558.
- (124) Malsch, K.; Rebentisch, R.; Swiderek, P.; Hohlneicher, G. Excited States of Acetylene: A CASPT2 Study. *Theor. Chem. Acc.* **1998**, *100*, 171–182.
- (125) Zyubin, A. S.; Mebel, A. M. Accurate Prediction of Excitation Energies to High-Lying Rydberg Electronic States: Rydberg States of Acetylene as a Case Study. *J. Chem. Phys.* **2003**, *119*, 6581–6587.
- (126) Ventura, E.; Dallos, M.; Lischka, H. The Valence-Excited States T₁–T₄ and S₁–S₂ of Acetylene: A High-Level MR-CISD and MR-AQCC Investigation of Stationary Points, Potential Energy Surfaces, and Surface Crossings. *J. Chem. Phys.* **2003**, *118*, 1702–1713.
- (127) Angeli, C.; Borini, S.; Ferrighi, L.; Cimraglia, R. Ab Initio N-Electron Valence State Perturbation Theory Study of the Adiabatic Transitions in Carbonyl Molecules: Formaldehyde, Acetaldehyde, and Acetone. *J. Chem. Phys.* **2005**, *122*, 114304.
- (128) Send, R.; Valsson, O.; Filippi, C. Electronic Excitations of Simple Cyanine Dyes: Reconciling Density Functional and Wave Function Methods. *J. Chem. Theory Comput.* **2011**, *7*, 444–455.
- (129) Dressler, R.; Allan, M. A Dissociative Electron Attachment, Electron Transmission, and Electron EnergyLoss Study of the Temporary Negative Ion of Acetylene. *J. Chem. Phys.* **1987**, *87*, 4510–4518.
- (130) Serrano-Andrés, L.; Merchán, M.; Nebot-Gil, I.; Lindh, R.; Roos, B. O. Towards an Accurate Molecular Orbital Theory for Excited States: Ethene, Butadiene, and Hexatriene. *J. Chem. Phys.* **1993**, *98*, 3151–3162.
- (131) Feller, D.; Peterson, K. A.; Davidson, E. R. A Systematic Approach to Vertically Excited States of Ethylene Using Configuration Interaction and Coupled Cluster Techniques. *J. Chem. Phys.* **2014**, *141*, 104302.
- (132) Shen, J.; Li, S. Block Correlated Coupled Cluster Method with the Complete Active-Space Self-Consistent-Field Reference Function: Applications for Low-Lying Electronic Excited States. *J. Chem. Phys.* **2009**, *131*, 174101.
- (133) Angeli, C. On the Nature of the Ionic Excited States: The V State of Ethene as a Prototype. *J. Comput. Chem.* **2009**, *30*, 1319–1333.
- (134) Müller, T.; Lischka, H. Simultaneous Calculation of Rydberg and Valence Excited States of Formaldehyde. *Theor. Chem. Acc.* **2001**, *106*, 369–378.
- (135) Schautz, F.; Buda, F.; Filippi, C. Excitations in Photoactive Molecules from Quantum Monte Carlo. *J. Chem. Phys.* **2004**, *121*, 5836–5844.
- (136) Paone, S.; Moule, D.; Bruno, A.; Steer, R. Vibronic Analyses of the Rydberg and Lower Intravalence Electronic Transitions in Thioacetone. *J. Mol. Spectrosc.* **1984**, *107*, 1–11.
- (137) Foresman, J. B.; Head-Gordon, M.; Pople, J. A.; Frisch, M. J. Toward a Systematic Molecular Orbital Theory for Excited States. *J. Phys. Chem.* **1992**, *96*, 135–149.
- (138) Hadad, C. M.; Foresman, J. B.; Wiberg, K. B. Excited States of Carbonyl Compounds. 1. Formaldehyde and Acetaldehyde. *J. Phys. Chem.* **1993**, *97*, 4293–4312.
- (139) Gwaltney, S. R.; Bartlett, R. J. An Application of the Equation-Of-Motion Coupled Cluster Method to the Excited States of Formaldehyde, Acetaldehyde, and Acetone. *Chem. Phys. Lett.* **1995**, *241*, 26–32.
- (140) Wiberg, K. B.; Stratmann, R. E.; Frisch, M. J. A Time-Dependent Density Functional Theory Study of the Electronically Excited States of Formaldehyde, Acetaldehyde and Acetone. *Chem. Phys. Lett.* **1998**, *297*, 60–64.
- (141) Wiberg, K. B.; de Oliveira, A. E.; Trucks, G. A Comparison of the Electronic Transition Energies for Ethene, Isobutene, Formaldehyde, and Acetone Calculated Using RPA, TDDFT, and EOM-CCSD. Effect of Basis Sets. *J. Phys. Chem. A* **2002**, *106*, 4192–4199.

- (142) Li, X.; Paldus, J. Multi-Reference State-Universal Coupled-Cluster Approaches to Electronically Excited States. *J. Chem. Phys.* **2011**, *134*, 214118.
- (143) Furche, F.; Ahlrichs, R. Adiabatic Time-Dependent Density Functional Methods for Excited States Properties. *J. Chem. Phys.* **2002**, *117*, 7433–7447.
- (144) Hättig, C. Response Theory and Molecular Properties (A Tribute to Jan Lindenberg and Poul Jørgensen). *Adv. Quantum Chem.* **2005**, *50*, 37–60.
- (145) Walzl, K. N.; Koerting, C. F.; Kuppermann, A. Electron-Impact Spectroscopy of Acetaldehyde. *J. Chem. Phys.* **1987**, *87*, 3796–3803.
- (146) Robin, M. B.; Basch, H.; Kuebler, N. A.; Wiberg, K. B.; Ellison, G. B. Optical Spectra of Small Rings. II. The Unsaturated Three-Membered Rings. *J. Chem. Phys.* **1969**, *51*, 45–52.
- (147) Sauer, I.; Grezzo, L. A.; Staley, S. W.; Moore, J. H. Low-Energy Singlet-Triplet and Singlet-Singlet Transitions in Cycloalkenes. *J. Am. Chem. Soc.* **1976**, *98*, 4218–4222.
- (148) McGlynn, S. P.; Rabalais, J. W.; McDonald, J. R.; Scherr, V. M. Electronic Spectroscopy of Isoelectronic Molecules. II. Linear Triatomic Groupings Containing Sixteen Valence Electrons. *Chem. Rev.* **1971**, *71*, 73–108.
- (149) Fedorov, I.; Koziol, L.; Li, G.; Parr, J. A.; Krylov, A. I.; Reisler, H. Theoretical and Experimental Investigations of the Electronic Rydberg States of Diazomethane: Assignments and State Interactions. *J. Phys. Chem. A* **2007**, *111*, 4557–4566.
- (150) Rittby, M.; Pal, S.; Bartlett, R. J. Multireference Coupled-Cluster Method: Ionization Potentials and Excitation Energies for Ketene and Diazomethane. *J. Chem. Phys.* **1989**, *90*, 3214–3220.
- (151) Gingell, J.; Mason, N.; Zhao, H.; Walker, I.; Siggel, M. Vuv Optical-Absorption and Electron-Energy-Loss Spectroscopy of Formamide. *Chem. Phys.* **1997**, *220*, 191–205.
- (152) Frueholz, R. P.; Flicker, W. M.; Kuppermann, A. Excited Electronic States of Ketene. *Chem. Phys. Lett.* **1976**, *38*, 57–60.
- (153) Xiao, H.; Maeda, S.; Morokuma, K. CASPT2 Study of Photodissociation Pathways of Ketene. *J. Phys. Chem. A* **2013**, *117*, 7001–7008.
- (154) Nooijen, M. First-Principles Simulation of the UV Absorption Spectrum of Ketene. *Int. J. Quantum Chem.* **2003**, *95*, 768–783.
- (155) Dixon, R. N.; Kroto, H. W. The Electronic Spectrum of Nitrosomethane, CH₃NO. *Proc. R. Soc. London, Ser. A* **1965**, *283*, 423–432.
- (156) Arenas, J. F.; Otero, J. C.; Peláez, D.; Soto, J. CASPT2 Study of the Decomposition of Nitrosomethane and Its Tautomerization Reactions in the Ground and Low-Lying Excited States. *J. Org. Chem.* **2006**, *71*, 983–991.
- (157) Dolgov, E. K.; Bataev, V. A.; Pupyshev, V. I.; Godunov, I. A. Ab Initio Description of the Structure and Dynamics of the Nitrosomethane Molecule in the First Excited Singlet and Triplet Electronic States. *Int. J. Quantum Chem.* **2004**, *96*, 589–597.
- (158) Reisler, H.; Krylov, A. I. Interacting Rydberg and Valence States in Radicals and Molecules: Experimental and Theoretical Studies. *Int. Rev. Phys. Chem.* **2009**, *28*, 267–308.
- (159) Jacquemin, D.; Duchemin, I.; Blase, X. Is the Bethe–Salpeter Formalism Accurate for Excitation Energies? Comparisons with TD-DFT, CASPT2, and EOM-CCSD. *J. Phys. Chem. Lett.* **2017**, *8*, 1524–1529.
- (160) Coe, J. P.; Paterson, M. J. State-Averaged Monte Carlo Configuration Interaction Applied to Electronically Excited States. *J. Chem. Phys.* **2013**, *139*, 154103.
- (161) Habas, M. P.; Dargelos, A. Ab Initio CI Calculations of Electronic and Vibrational Spectra of Diazomethane. *Chem. Phys.* **1995**, *199*, 177–182.
- (162) Serrano-Andrés, L.; Fülischer, M. P. Theoretical Study of the Electronic Spectroscopy of Peptides. I. The Peptidic Bond: Primary, Secondary, and Tertiary Amides. *J. Am. Chem. Soc.* **1996**, *118*, 12190–12199.
- (163) Besley, N. A.; Hirst, J. D. Ab Initio Study of the Electronic Spectrum of Formamide with Explicit Solvent. *J. Am. Chem. Soc.* **1999**, *121*, 8559–8566.
- (164) Szalay, P. G.; Császár, A. G.; Nemes, L. Electronic States of Ketene. *J. Chem. Phys.* **1996**, *105*, 1034–1045.
- (165) Lacombe, S.; Loudet, M.; Dargelos, A.; Camou, J. Calculation of the Electronic and Photoelectronic Spectra of Nitroso Compounds: A Reinvestigation by Use of Configuration Interaction Methods. *Chem. Phys.* **2000**, *258*, 1–12.
- (166) Dolgov, E. K.; Bataev, V. A.; Godunov, I. A. Structure of the Nitrosomethane Molecule (CH₃NO) in the Ground Electronic State: Testing of Ab Initio Methods for the Description of Potential Energy Surface. *Int. J. Quantum Chem.* **2004**, *96*, 193–201.
- (167) Moore, B., II; Autschbach, J. Longest-Wavelength Electronic Excitations of Linear Cyanines: The Role of Electron Delocalization and of Approximations in Time-Dependent Density Functional Theory. *J. Chem. Theory Comput.* **2013**, *9*, 4991–5003.
- (168) Boulanger, P.; Jacquemin, D.; Duchemin, I.; Blase, X. Fast and Accurate Electronic Excitations in Cyanines with the Many-Body Bethe-Salpeter Approach. *J. Chem. Theory Comput.* **2014**, *10*, 1212–1218.
- (169) Zhekova, H.; Krykunov, M.; Autschbach, J.; Ziegler, T. Applications of Time Dependent and Time Independent Density Functional Theory to the First π to π^* Transition in Cyanine Dyes. *J. Chem. Theory Comput.* **2014**, *10*, 3299–3307.
- (170) Le Guennic, B.; Jacquemin, D. Taking Up the Cyanine Challenge with Quantum Tools. *Acc. Chem. Res.* **2015**, *48*, 530–537.
- (171) Tarte, P. Recherches Spectroscopiques sur les Composés Nitrosés. *Bull. Soc. Chim. Belg.* **1954**, *63*, 525–541.
- (172) Ernsting, N. P.; Pfab, J.; Romelt, J. Geometry Changes Accompanying Electronic Excitation of Nitrosomethane in the 650 nm Region. *J. Chem. Soc., Faraday Trans. 2* **1978**, *74*, 2286–2294.
- (173) Gordon, R. D.; Luck, P. Conformational Changes Accompanying Electronic Excitation of CD₃NO. *Chem. Phys. Lett.* **1979**, *65*, 480–483.
- (174) Jacquemin, D. What is the Key for Accurate Absorption and Emission Calculations? Energy or Geometry? *J. Chem. Theory Comput.* **2018**, *14*, 1534–1543.
- (175) The larger deviations with CCSDTQ is likely due to the larger number of Rydberg states, which are more basis set sensitive, in that set. Let us note also that several CCSDTQ/aug-cc-pVTZ reference values are obtained by correcting CCSDTQ/aug-cc-pVDZ values. However, such a procedure has been shown to be very robust above, so that it does not impact the present analysis.